

References

- [1] Gervasi J, Dubois L, Thomas D. Screening tests of new hybrid solvents for the post-combustion CO₂ capture process by chemical absorption. *Energy Procedia* 2014;63:1854-62. <https://doi.org/10.1016/j.egypro.2014.11.193>
- [2] Wilberforce T, Baroutaji A, Soudan B, Al-Alami AH, Olabi AG. Outlook of carbon capture technology and challenges. *Sci Total Environ* 2019;657:56-72. <https://doi.org/10.1016/j.scitotenv.2018.11.424>
- [3] Shen Y, Jiang C, Zhang S, Chen J, Wang L, Chen J. Biphasic solvent for CO₂ capture: Amine property-performance and heat duty relationship. *Appl Energy* 2018;230:726-33. <https://doi.org/10.1016/j.apenergy.2018.09.005>
- [4] Wai SK, Saiwan C, Idem R, Supap T, Nwaoha C. Carbon Dioxide (CO₂) Solubility in Diethylenetriamine and 2-Amino-2-Methyl-1-Propanol (DETA–AMP) Solvent System for Amine–Based CO₂ Capture in Flue Gas from Coal Combustion. *Energy Procedia* 2017;114:1973-9. <https://doi.org/10.1016/j.egypro.2017.03.1329>
- [5] Singh S, Pandey D, Mondal MK. New Experimental Data on Equilibrium CO₂ Loading into Aqueous 3-Dimethyl Amino-1-propanol and 1, 5-Diamino-2-methylpentane Blend: Empirical Model and CO₂ Absorption Enthalpy. *J Chem Eng Data* 2020;66(1):740-8. <https://doi.org/10.1021/acs.jced.0c00851>
- [6] Gautam A, Mondal MK. Post-combustion capture of CO₂ using novel aqueous Triethylenetetramine and 2-Dimethylaminoethanol amine blend: Equilibrium CO₂ loading-empirical model and optimization, CO₂ desorption, absorption heat, and ¹³C NMR analysis. *Fuel*. 2023;331:125864. <https://doi.org/10.1016/j.fuel.2022.125864>

- [7] Bougie F, Iliuta MC. Sterically hindered amine-based absorbents for the removal of CO₂ from gas streams. *J Chem Eng Data* 2012;57(3):635-69. <https://doi.org/10.1021/je200731v>
- [8] Udara SA, Kohilan R, Lakshan MA, Madalagama ML, Pathirana PP, Sandupama PS. Simulation of carbon dioxide capture for industrial applications. *Energy Rep* 2020;6:659-63. <https://doi.org/10.1016/j.egy.2019.11.134>
- [9] Bailera M, Lisbona P, Peña B, Romeo LM. A review on CO₂ mitigation in the Iron and Steel industry through Power to X processes. *J CO₂ Util.* 2021;46:101456. <https://doi.org/10.1016/j.jcou.2021.101456>
- [10] Bains P, Psarras P, Wilcox J. CO₂ capture from the industry sector. *Prog Energ Combust Sci* 2017;63:146-72. <https://doi.org/10.1016/j.pecs.2017.07.001>
- [11] Youn MH, Park KT, Lee YH, Kang SP, Lee SM, Kim SS, Kim YE, Ko YN, Jeong SK, Lee W. Carbon dioxide sequestration process for the cement industry. *J CO₂ Util* 2019;34:325-34. <https://doi.org/10.1016/j.jcou.2019.07.023>
- [12] Nazir SM, Cloete JH, Cloete S, Amini S. Efficient hydrogen production with CO₂ capture using gas switching reforming. *Energy* 2019;185:372-85. <https://doi.org/10.1016/j.energy.2019.07.072>
- [13] Aghel B, Behaein S, Alobiad F. CO₂ capture from biogas by biomass-based adsorbents: A review. *Fuel*. 2022;328:125276. <https://doi.org/10.1016/j.fuel.2022.125276>
- [14] Zhang J, Agar DW, Zhang X, Geuzebroek F. CO₂ absorption in biphasic solvents with enhanced low temperature solvent regeneration. *Energy Procedia* 2011;4:67-74. <https://doi.org/10.1016/j.egypro.2011.01.024>

- [15] Zhang J, Nwani O, Tan Y, Agar DW. Carbon dioxide absorption into biphasic amine solvent with solvent loss reduction. *Chem Eng Res Des* 2011;89(8):1190-6. <https://doi.org/10.1016/j.cherd.2011.02.005>
- [16] Zhang J, Qiao Y, Wang W, Misch R, Hussain K, Agar DW. Development of an energy-efficient CO₂ capture process using thermomorphic biphasic solvents. *Energy Procedia* 2013;37:1254-61. <https://doi.org/10.1016/j.egypro.2013.05.224>
- [17] Kumar S, Mondal MK. Equilibrium solubility of CO₂ in aqueous binary mixture of 2-(diethylamine) ethanol and 1, 6-hexamethyldiamine. *Korean J Chem Eng* 2018;35(6):1335-40. <https://doi.org/10.1007/s11814-018-0044-6>
- [18] Kumar S, Mondal MK. Equilibrium Solubility of CO₂ in Aqueous Blend of 2-(Diethylamine) ethanol and 2-(2-Aminoethylamine) ethanol. *J Chem Eng Data* 2018;63(5):1163-9. <https://doi.org/10.1021/acs.jced.7b00544>
- [19] Borhani TN, Wang M. Role of solvents in CO₂ capture processes: The review of selection and design methods. *Renew Sustain Energy Rev.* 2019;114:109299. <https://doi.org/10.1016/j.rser.2019.109299>
- [20] Nwaoha C, Saiwan C, Tontiwachwuthikul P, Supap T, Rongwong W, Idem R, AL-Marri MJ, Benamor A. Carbon dioxide (CO₂) capture: Absorption-desorption capabilities of 2-amino-2-methyl-1-propanol (AMP), piperazine (PZ) and monoethanolamine (MEA) tri-solvent blends. *J Nat Gas Sci Eng* 2016;33:742-50. <https://doi.org/10.1016/j.jngse.2016.06.002>
- [21] Wang M, Rahimi M, Kumar A, Hariharan S, Choi W, Hatton TA. Flue gas CO₂ capture via electrochemically mediated amine regeneration: System design and

- performance. Appl Energy. 2019;255:113879.
<https://doi.org/10.1016/j.apenergy.2019.113879>
- [22] Pinto DD, Zaidy SA, Hartono A, Svendsen HF. Evaluation of a phase change solvent for CO₂ capture: Absorption and desorption tests. Int J Greenh Gas Control 2014;28:318-27. <https://doi.org/10.1016/j.ijggc.2014.07.002>
- [23] Theo WL, Lim JS, Hashim H, Mustaffa AA, Ho WS. Review of pre-combustion capture and ionic liquid in carbon capture and storage. Appl Energy 2016;183:1633-63. <https://doi.org/10.1016/j.apenergy.2016.09.103>
- [24] Leung DY, Caramanna G, Maroto-Valer MM. An overview of current status of carbon dioxide capture and storage technologies. Renew Sustain Energy Rev 2014;39:426-43. <https://doi.org/10.1016/j.rser.2014.07.093>
- [25] Fu D, Xie J. Absorption capacity and viscosity for CO₂ capture process using [N1111][Gly] promoted K₂CO₃ aqueous solution. J Chem Thermodyn 2016;102:310-5. <https://doi.org/10.1016/j.jct.2016.07.026>
- [26] Song C, Liu Q, Deng S, Li H, Kitamura Y. Cryogenic-based CO₂ capture technologies: State-of-the-art developments and current challenges. Renew Sustain Energy Rev 2019;101:265-78. <https://doi.org/10.1016/j.rser.2018.11.018>
- [27] Siagian UW, Raksajati A, Himma NF, Khoiruddin K, Wenten IG. Membrane-based carbon capture technologies: Membrane gas separation vs. membrane contactor. J Nat Gas Sci Eng 2019;67:172-95. <https://doi.org/10.1016/j.jngse.2019.04.008>
- [28] Olajire AA. CO₂ capture and separation technologies for end-of-pipe applications—a review. Energy 2010;35(6):2610-28. <https://doi.org/10.1016/j.energy.2010.02.030>

- [29] Ramdin M, de Loos TW, Vlucht TJ. State-of-the-art of CO₂ capture with ionic liquids. *Ind Eng Chem Res* 2012;51(24):8149-77. <https://doi.org/10.1021/ie3003705>
- [30] Zhicheng X, Shujuan W, Bo Z, Changhe C. Study on potential biphasic solvents: Absorption capacity, CO₂ loading and reaction rate. *Energy Procedia* 2013;37:494-8. <https://doi.org/10.1016/j.egypro.2013.05.135>
- [31] Mukhtar A, Saqib S, Mellon NB, Babar M, Rafiq S, Ullah S, Bustam MA, Al-Sehemi AG, Muhammad N, Chawla M. CO₂ capturing, thermo-kinetic principles, synthesis and amine functionalization of covalent organic polymers for CO₂ separation from natural gas: A review. *J Nat Gas Sci Eng.* 2020;77:103203. <https://doi.org/10.1016/j.jngse.2020.103203>
- [32] Kanniche M, Gros-Bonnivard R, Jaud P, Valle-Marcos J, Amann JM, Bouallou C. Pre-combustion, post-combustion and oxy-combustion in thermal power plant for CO₂ capture. *Appl Therm Eng* 2010;30(1):53-62. <https://doi.org/10.1016/j.applthermaleng.2009.05.005>
- [33] Lee JH, Im K, Han S, Yoo SJ, Kim J, Kim JH. Bimodal-porous hollow MgO sphere embedded mixed matrix membranes for CO₂ capture. *Sep Purif Technol.* 2020;250:117065. <https://doi.org/10.1016/j.seppur.2020.117065>
- [34] Liu B, Zhang M, Yang X, Wang T. Simulation and energy analysis of CO₂ capture from CO₂-EOR extraction gas using cryogenic fractionation. *J Taiwan Inst Chem Eng* 2019;103:67-74. <https://doi.org/10.1016/j.jtice.2019.07.008>
- [35] Araújo OD, de Medeiros JL. Carbon capture and storage technologies: present scenario and drivers of innovation. *Curr Opin Chem Eng* 2017;17:22-34. <https://doi.org/10.1016/j.coche.2017.05.004>

- [36] Hu X, Huang J, He X, Luo Q, Li CE, Zhou C, Zhang R. Analyzing the potential benefits of trio-amine systems for enhancing the CO₂ desorption processes. *Fuel*. 2022;316:123216. <https://doi.org/10.1016/j.fuel.2022.123216>
- [37] Zhou X, Liu F, Lv B, Zhou Z, Jing G. Evaluation of the novel biphasic solvents for CO₂ capture: Performance and mechanism. *Int J Greenh Gas Control* 2017;60:120-8. <https://doi.org/10.1016/j.ijggc.2017.03.013>
- [38] Ye Q, Wang X, Lu Y. Screening and evaluation of novel biphasic solvents for energy-efficient post-combustion CO₂ capture. *Int J Greenh Gas Control* 2015;39:205-14. <https://doi.org/10.1016/j.ijggc.2015.05.025>
- [39] Zhuang Q, Clements B, Dai J, Carrigan L. Ten years of research on phase separation absorbents for carbon capture: Achievements and next steps. *Int J Greenh Gas Control* 2016;52:449-60. <https://doi.org/10.1016/j.ijggc.2016.04.022>
- [40] Kim YE, Park JH, Yun SH, Nam SC, Jeong SK, Yoon YI. Carbon dioxide absorption using a phase transitional alkanolamine – alcohol mixture. *J Ind Eng Chem* 2014;20(4):1486-92. <https://doi.org/10.1016/j.jiec.2013.07.036>
- [41] Wang R, Liu S, Li Q, Zhang S, Wang L, An S. CO₂ capture performance and mechanism of blended amine solvents regulated by N-methylcyclohexylamine. *Energy*. 2021;215:119209. <https://doi.org/10.1016/j.energy.2020.119209>
- [42] Liu F, Fang M, Dong W, Wang T, Xia Z, Wang Q, Luo Z. Carbon dioxide absorption in aqueous alkanolamine blends for biphasic solvents screening and evaluation. *Appl Energy* 2019;233:468-77. <https://doi.org/10.1016/j.apenergy.2018.10.007>

- [43] Zhang J, Qiao Y, Agar DW. Improvement of lipophilic-amine-based thermomorphic biphasic solvent for energy-efficient carbon capture. *Energy Procedia* 2012;23:92-101. <https://doi.org/10.1016/j.egypro.2012.06.072>
- [44] Xu Z, Wang S, Chen C. CO₂ absorption by biphasic solvents: Mixtures of 1, 4-Butanediamine and 2-(Diethylamino)-ethanol. *Int J Greenh Gas Control* 2013;16:107-15. <https://doi.org/10.1016/j.ijggc.2013.03.013>
- [45] Ye Q, Lu H, Du Y, Zhang S, Wang X, Lu Y. Experimental investigation of the absorption, phase transition, and desorption behavior of biphasic solvent blends for post combustion CO₂ capture. *Energy Procedia* 2017;114:813-22. <https://doi.org/10.1016/j.egypro.2017.03.1223>
- [46] Machida H, Oba K, Tomikawa T, Esaki T, Yamaguchi T, Horizoe H. Development of phase separation solvent for CO₂ capture by aqueous (amine + ether) solution. *J Chem Thermodyn* 2017;113:64-70. <https://doi.org/10.1016/j.jct.2017.05.043>
- [47] Shen Y, Chen H, Wang J, Zhang S, Jiang C, Ye J, Wang L, Chen J. Two-stage interaction performance of CO₂ absorption into biphasic solvents: Mechanism analysis, quantum calculation and energy consumption. *Appl Energy*. 2020;260:114343. <https://doi.org/10.1016/j.apenergy.2019.114343>
- [48] White CM, Strazisar BR, Granite EJ, Hoffman JS, Pennline HW. Separation and capture of CO₂ from large stationary sources and sequestration in geological formations—coal beds and deep saline aquifers. *J Air Waste Manage* 2003;53(6):645-715. <https://doi.org/10.1080/10473289.2003.10466206>

- [49] Mondal MK, Balsora HK, Varshney P. Progress and trends in CO₂ capture/separation technologies: A review. *Energy* 2012;46(1):431-41. <https://doi.org/10.1016/j.energy.2012.08.006>
- [50] Lee HJ, Lee JD, Linga P, Englezos P, Kim YS, Lee MS, Do Kim Y. Gas hydrate formation process for pre-combustion capture of carbon dioxide. *Energy* 2010;35(6):2729-33. <https://doi.org/10.1016/j.energy.2009.05.026>
- [51] Wu F, Argyle MD, Dellenback PA, Fan M. Progress in O₂ separation for oxy-fuel combustion—A promising way for cost-effective CO₂ capture: A review. *Prog Energ Combust Sci* 2018;67:188-205. <https://doi.org/10.1016/j.peccs.2018.01.004>
- [52] Li K, Feron PH, Jones TW, Jiang K, Bennett RD, Hollenkamp AF. Energy harvesting from amine-based CO₂ capture: proof-of-concept based on mono-ethanolamine. *Fuel*. 2020;263:116661. <https://doi.org/10.1016/j.fuel.2019.116661>
- [53] Wang R, Zhao H, Wang Y, Qi C, Zhang S, Wang L, Li M. Development of biphasic solvent for CO₂ capture by tailoring the polarity of amine solution. *Fuel*. 2022;325:124885. <https://doi.org/10.1016/j.fuel.2022.124885>
- [54] Rubin ES, Mantripragada H, Marks A, Versteeg P, Kitchin J. The outlook for improved carbon capture technology. *Prog Energ Combust Sci* 2012;38(5):630-71. <https://doi.org/10.1016/j.peccs.2012.03.003>
- [55] Bounaceur R, Lape N, Roizard D, Vallieres C, Favre E. Membrane processes for post-combustion carbon dioxide capture: a parametric study. *Energy* 2006;31(14):2556-70. <https://doi.org/10.1016/j.energy.2005.10.038>
- [56] Chao C, Deng Y, Dewil R, Baeyens J, Fan X. Post-combustion carbon capture. *Renew Sustain Energy Rev*. 2021;138:110490. <https://doi.org/10.1016/j.rser.2020.110490>

- [57] Sreedhar I, Nahar T, Venugopal A, Srinivas B. Carbon capture by absorption—Path covered and ahead. *Renew Sustain Energy Rev* 2017;76:1080-107. <https://doi.org/10.1016/j.rser.2017.03.109>
- [58] Bhowan AS, Freeman BC. Analysis and status of post-combustion carbon dioxide capture technologies. *Environ Sci Technol* 2011;45(20):8624-32. <https://doi.org/10.1021/es104291d>
- [59] Metz B, Davidson O, De Coninck HC, Loos M, Meyer L. IPCC special report on carbon dioxide capture and storage. Cambridge: Cambridge University Press; 2005.
- [60] Wienchol P, Szlęk A, Ditaranto M. Waste-to-energy technology integrated with carbon capture—Challenges and opportunities. *Energy*. 2020;198:117352. <https://doi.org/10.1016/j.energy.2020.117352>
- [61] Pandey D, Mondal MK. Equilibrium CO₂ solubility in the aqueous mixture of MAE and AEEA: Experimental study and development of modified thermodynamic model. *Fluid Ph Equilibria*. 2020;522:112766. <https://doi.org/10.1016/j.fluid.2020.112766>
- [62] Chen WH, Chen SM, Hung CI. Carbon dioxide capture by single droplet using Selexol, Rectisol and water as absorbents: A theoretical approach. *Appl Energy* 2013;111:731-41. <https://doi.org/10.1016/j.apenergy.2013.05.051>
- [63] Park SH, Lee SJ, Lee JW, Chun SN, Lee JB. The quantitative evaluation of two-stage pre-combustion CO₂ capture processes using the physical solvents with various design parameters. *Energy* 2015;81:47-55. <https://doi.org/10.1016/j.energy.2014.10.055>
- [64] Kapetaki Z, Brandani P, Brandani S, Ahn H. Process simulation of a dual-stage Selexol process for 95% carbon capture efficiency at an integrated gasification combined

- cycle power plant. *Int J Greenh Gas Control* 2015;39:17-26.
<https://doi.org/10.1016/j.ijggc.2015.04.015>
- [65] Wang L, Zhang Y, Wang R, Li Q, Zhang S, Li M, Liu J, Chen B. Advanced monoethanolamine absorption using sulfolane as a phase splitter for CO₂ capture. *Environ Sci Technol* 2018;52(24):14556-63. <https://doi.org/10.1021/acs.est.8b05654>
- [66] Wang L, Yu S, Li Q, Zhang Y, An S, Zhang S. Performance of sulfolane/DETA hybrids for CO₂ absorption: Phase splitting behavior, kinetics and thermodynamics. *Appl Energy* 2018;228:568-76. <https://doi.org/10.1016/j.apenergy.2018.06.077>
- [67] Wappel D, Gronald G, Kalb R, Draxler J. Ionic liquids for post-combustion CO₂ absorption. *Int J Greenh Gas Control* 2010;4(3):486-94.
<https://doi.org/10.1016/j.ijggc.2009.11.012>
- [68] Sistla YS, Khanna A. CO₂ absorption studies in amino acid-anion based ionic liquids. *Chem Eng J* 2015;273:268-76. <https://doi.org/10.1016/j.cej.2014.09.043>
- [69] Jilvero H, Normann F, Andersson K, Johnsson F. The Rate of CO₂ Absorption in Ammonia-Implications on Absorber Design. *Ind Eng Chem Res* 2014;53(16):6750-8.
<https://doi.org/10.1021/ie403346a>
- [70] Dai Z, Deng L. Membrane absorption using ionic liquid for pre-combustion CO₂ capture at elevated pressure and temperature. *Int J Greenh Gas Control* 2016;54:59-69.
<https://doi.org/10.1016/j.ijggc.2016.09.001>
- [71] Oexmann J, Kather A. Minimising the regeneration heat duty of post-combustion CO₂ capture by wet chemical absorption: The misguided focus on low heat of absorption solvents. *Int J Greenh Gas Control* 2010;4(1):36-43.
<https://doi.org/10.1016/j.ijggc.2009.09.010>

- [72] Nimmanterdwong P, Chalermisinsuwan B, Piumsomboon P. Emergy analysis of three alternative carbon dioxide capture processes. *Energy* 2017;128:101-8. <https://doi.org/10.1016/j.energy.2017.03.154>
- [73] Kumar S, Padhan R, Mondal MK. Equilibrium solubility measurement and modeling of CO₂ absorption in aqueous blend of 2-(diethyl amino) ethanol and ethylenediamine. *J Chem Eng Data* 2020;65(2):523-31. <https://doi.org/10.1021/acs.jced.9b00699>
- [74] Guo H, Li C, Shi X, Li H, Shen S. Nonaqueous amine-based absorbents for energy efficient CO₂ capture. *Appl Energy* 2019;239:725-34. <https://doi.org/10.1016/j.apenergy.2019.02.019>
- [75] Lampreia IM, Santos ÂF, Moita ML, Nobre LC. Thermodynamic study of aqueous 2-(isopropylamino) ethanol. A sterically hindered new amine absorbent for CO₂ capture. *J Chem Thermodyn* 2015;81:167-76. <https://doi.org/10.1016/j.jct.2014.10.004>
- [76] Idris Z, Eimer DA. Representation of CO₂ absorption in sterically hindered amines. *Energy Procedia* 2014;51:247-52. <https://doi.org/10.1016/j.egypro.2014.07.028>
- [77] Freeman SA, Dugas R, Van Wagener DH, Nguyen T, Rochelle GT. Carbon dioxide capture with concentrated, aqueous piperazine. *Int J Greenh Gas Control* 2010;4(2):119-24. <https://doi.org/10.1016/j.ijggc.2009.10.008>
- [78] Arshad MW, von Solms N, Thomsen K. Thermodynamic modeling of liquid–liquid phase change solvents for CO₂ capture. *Int J Greenh Gas Control* 2016;53:401-24. <https://doi.org/10.1016/j.ijggc.2016.08.014>
- [79] Caplow M. Kinetics of carbamate formation and breakdown. *J Am Chem Soc* 1968;90(24):6795-803. <https://doi.org/10.1021/ja01026a041>

- [80] Danckwerts PV. The reaction of CO₂ with ethanolamines. *Chem Eng Sci* 1979;34(4):443-6. [https://doi.org/10.1016/0009-2509\(79\)85087-3](https://doi.org/10.1016/0009-2509(79)85087-3)
- [81] Crooks JE, Donnellan JP. Kinetics and mechanism of the reaction between carbon dioxide and amines in aqueous solution. *J Chem Soc Perkin Trans 2* 1989:331-3. <https://doi.org/10.1039/P29890000331>
- [82] Da Silva EF, Svendsen HF. Ab initio study of the reaction of carbamate formation from CO₂ and alkanolamines. *Ind Eng Chem Res* 2004;43(13):3413-8. <https://doi.org/10.1021/ie030619k>
- [83] Ben-Mansour R, Habib MA, Bamidele OE, Basha M, Qasem NA, Peedikakkal A, Laoui T, Ali MJ. Carbon capture by physical adsorption: materials, experimental investigations and numerical modeling and simulations—a review. *Appl Energy* 2016;161:225-55. <https://doi.org/10.1016/j.apenergy.2015.10.011>
- [84] Agnihotri N, Mondal MK. Catalytic pyrolysis for upgrading of biooil obtained from biomass. *Biofuels and Bioenergy*: Elsevier; 2022.
- [85] May-Britt HÄ, Kim TJ, Li B, inventors; Ntnu Technology Transfer As, assignee, 2005. Membrane for separating CO₂ and process for the production thereof. WO2005089907A1.
- [86] Ahmad NN, Leo CP, Ahmad AL. Effects of solvent and ionic liquid properties on ionic liquid enhanced polysulfone/SAPO-34 mixed matrix membrane for CO₂ removal. *Micropor Mesopor Mat* 2019;283:64-72. <https://doi.org/10.1016/j.micromeso.2019.04.001>
- [87] Ahmad AL, Sunarti AR, Lee KT, Fernando WJ. CO₂ removal using membrane gas absorption. *Int J Greenh Gas Control* 2010;4(3):495-8. <https://doi.org/10.1016/j.ijggc.2009.12.003>

- [88] Yousef AM, El-Maghlany WM, Eldrainy YA, Attia A. New approach for biogas purification using cryogenic separation and distillation process for CO₂ capture. *Energy* 2018;156:328-51. <https://doi.org/10.1016/j.energy.2018.05.106>
- [89] Alami AH, Alasad S, Ali M, Alshamsi M. Investigating algae for CO₂ capture and accumulation and simultaneous production of biomass for biodiesel production. *Sci Total Environ.* 2021;759:143529. <https://doi.org/10.1016/j.scitotenv.2020.143529>
- [90] Klinthong W, Yang YH, Huang CH, Tan CS. A review: microalgae and their applications in CO₂ capture and renewable energy. *Aerosol Air Qual Res* 2015;15(2):712-42. <https://doi.org/10.4209/aaqr.2014.11.0299>
- [91] Daneshvar E, Wicker RJ, Show PL, Bhatnagar A. Biologically-mediated carbon capture and utilization by microalgae towards sustainable CO₂ biofixation and biomass valorization—A review. *Chem Eng J.* 2022;427:130884. <https://doi.org/10.1016/j.cej.2021.130884>
- [92] Huang D, Li MJ, Wang RL, Yang YW, Tao WQ. Advanced carbon sequestration by the hybrid system of photo bioreactor and microbial fuel cell with novel photocatalytic porous framework. *Bioresour Technol.* 2021;333:125182. <https://doi.org/10.1016/j.biortech.2021.125182>
- [93] Molino A, Mehariya S, Karatza D, Chianese S, Iovine A, Casella P, Marino T, Musmarra D. Bench-scale cultivation of microalgae *Scenedesmus almeriensis* for CO₂ capture and lutein production. *Energies* 2019;12(14):2806. <https://doi.org/10.3390/en12142806>

- [94] Zhang S, Shen Y, Wang L, Chen J, Lu Y. Phase change solvents for post-combustion CO₂ capture: Principle, advances, and challenges. *Appl Energy* 2019;239:876-97. <https://doi.org/10.1016/j.apenergy.2019.01.242>
- [95] Arshad MW, Fosbøl PL, von Solms N, Svendsen HF, Thomsen K. Heat of absorption of CO₂ in phase change solvents: 2-(diethylamino) ethanol and 3-(methylamino) propylamine. *J Chem Eng Data* 2013;58(7):1974-88. <https://doi.org/10.1021/je400289v>
- [96] Raynal L, Alix P, Bouillon PA, Gomez A, de Nailly ML, Jacquin M, Kittel J, Di Lella A, Mougin P, Trapy J. The DMX™ process: an original solution for lowering the cost of post-combustion carbon capture. *Energy Procedia* 2011;4:779-86. <https://doi.org/10.1016/j.egypro.2011.01.119>
- [97] Kumar S, Mondal MK. Selection of efficient absorbent for CO₂ capture from gases containing low CO₂. *Korean J Chem Eng* 2020;37(2):231-9. <https://doi.org/10.1007/s11814-019-0440-6>
- [98] Wang X, Akhmedov NG, Hopkinson D, Hoffman J, Duan Y, Egbebi A, Resnik K, Li B. Phase change amino acid salt separates into CO₂-rich and CO₂-lean phases upon interacting with CO₂. *Appl Energy* 2016;161:41-7. <https://doi.org/10.1016/j.apenergy.2015.09.094>
- [99] Zhuang Q, Clements B. CO₂ capture by biphasic absorbent–absorption performance and VLE characteristics. *Energy* 2018;147:169-76. <https://doi.org/10.1016/j.energy.2018.01.004>
- [100] Zhan X, Lv B, Yang K, Jing G, Zhou Z. Dual-functionalized ionic liquid biphasic solvent for carbon dioxide capture: high-efficiency and energy saving. *Environ Sci Technol* 2020;54(10):6281-8. <https://doi.org/10.1021/acs.est.0c00335>

- [101] Zarogiannis T, Papadopoulos AI, Seferlis P. Efficient selection of conventional and phase-change CO₂ capture solvents and mixtures based on process economic and operating criteria. *J Clean Prod.* 2020;272:122764. <https://doi.org/10.1016/j.jclepro.2020.122764>
- [102] Xu M, Wang S, Xu L. Screening of physical-chemical biphasic solvents for CO₂ absorption. *Int J Greenh Gas Control* 2019;85:199-205. <https://doi.org/10.1016/j.ijggc.2019.03.015>
- [103] Wang R, Liu S, Wang L, Li Q, Zhang S, Chen B, Jiang L, Zhang Y. Superior energy-saving splitter in monoethanolamine-based biphasic solvents for CO₂ capture from coal-fired flue gas. *Appl Energy* 2019;242:302-10. <https://doi.org/10.1016/j.apenergy.2019.03.138>
- [104] Zhang S, Shen Y, Wang L, Chen J, Lu Y. Phase change solvents for post-combustion CO₂ capture: Principle, advances, and challenges. *Appl Energy* 2019;239:876-97. <https://doi.org/10.1016/j.apenergy.2019.01.242>
- [105] Lv J, Liu S, Ling H, Gao H, Olson W, Li Q, Bairq ZA, Liang Z. Development of a promising biphasic absorbent for post combustion CO₂ capture: Sulfolane + 2-(methylamino) ethanol + H₂O. *Ind Eng Chem Res* 2020;59(32):14496-506. <https://doi.org/10.1021/acs.iecr.0c02389>
- [106] Luo W, Guo D, Zheng J, Gao S, Chen J. CO₂ absorption using biphasic solvent: blends of diethylenetriamine, sulfolane, and water. *Int J Greenh Gas Control* 2016;53:141-8. <https://doi.org/10.1016/j.ijggc.2016.07.036>
- [107] Papadopoulos AI, Tzirakis F, Tsivintzelis I, Seferlis P. Phase-change solvents and processes for postcombustion CO₂ capture: a detailed review. *Ind Eng Chem Res* 2019;58(13):5088-111. <https://doi.org/10.1021/acs.iecr.8b06279>

- [108] Huang Q, Jing G, Zhou X, Lv B, Zhou Z. A novel biphasic solvent of amino-functionalized ionic liquid for CO₂ capture: High efficiency and regenerability. *J CO₂ Util* 2018;25:22-30. <https://doi.org/10.1016/j.jcou.2018.03.001>
- [109] Lee J, Hong YK, You JK. Phase separation characteristics in biphasic solvents based on mutually miscible amines for energy efficient CO₂ capture. *Korean J Chem Eng* 2017;34(6):1840-5. <https://doi.org/10.1007/s11814-017-0067-4>
- [110] Dreillard M, Broutin P, Briot P, Huard T, Lettat A. Application of the DMXTM CO₂ capture process in steel industry. *Energy Procedia* 2017;114:2573-89. <https://doi.org/10.1016/j.egypro.2017.03.1415>
- [111] Cheng J, Li Y, Hu L, Liu J, Zhou J, Cen K. Characterization of CO₂ absorption and carbamate precipitate in phase-change N-methyl-1,3-diaminopropane/ N,N-dimethylformamide solvent. *Energ Fuel* 2017;31(12):13972-8. <https://doi.org/10.1021/acs.energyfuels.7b02627>
- [112] Li Y, Cheng J, Hu L, Liu N, Zhou J, Cen K. Regulating crystal structures of EDA-carbamates in solid–liquid phase-changing CO₂ capture solutions. *Fuel* 2019;252:47-54. <https://doi.org/10.1016/j.fuel.2019.04.051>
- [113] Li H, Guo H, Shen S. Low-energy-consumption CO₂ capture by liquid–solid phase change absorption using water-lean blends of amino acid salts and 2-alkoxyethanols. *ACS Sustain Chem Eng* 2020;8(34):12956-67. <https://doi.org/10.1021/acssuschemeng.0c03525>
- [114] Tran KV, Ando R, Yamaguchi T, Machida H, Norinaga K. Carbon dioxide absorption heat in liquid–liquid and solid–liquid phase-change solvents using continuous calorimetry. *Ind Eng Chem Res* 2020;59(8):3475-84. <https://doi.org/10.1021/acs.iecr.9b04672>

- [115] Vega F, Baena-Moreno FM, Fernandez LM, Portillo E, Navarrete B, Zhang Z. Current status of CO₂ chemical absorption research applied to CCS: Towards full deployment at industrial scale. *Appl Energy*. 2020;260:114313. <https://doi.org/10.1016/j.apenergy.2019.114313>
- [116] Aleixo M, Prigent M, Gibert A, Porcheron F, Mokbel I, Jose J, Jacquin M. Physical and chemical properties of DMXTM solvents. *Energy Procedia* 2011;4:148-55. <https://doi.org/10.1016/j.egypro.2011.01.035>
- [117] Wang N, Peng Z, Gao H, Sema T, Shi J, Liang Z. New insight and evaluation of secondary Amine/N-butanol biphasic solutions for CO₂ Capture: Equilibrium Solubility, phase separation Behavior, absorption Rate, desorption Rate, energy consumption and ion species. *Chem Eng J*. 2022;431:133912. <https://doi.org/10.1016/j.cej.2021.133912>
- [118] Barzagli F, Mani F, Peruzzini M. Novel water-free biphasic absorbents for efficient CO₂ capture. *Int J Greenh Gas Control* 2017;60:100-9. <https://doi.org/10.1016/j.ijggc.2017.03.010>
- [119] Lockwood T. A comparative review of next-generation carbon capture technologies for coal-fired power plant. *Energy Procedia* 2017;114:2658-70. <https://doi.org/10.1016/j.egypro.2017.03.1850>
- [120] Li Q, Gao G, Wang R, Zhang S, An S, Wang L. Role of 1-methylimidazole in regulating the CO₂ capture performance of triethylenetetramine-based biphasic solvents. *Int J Greenh Gas Control*. 2021;108:103330. <https://doi.org/10.1016/j.ijggc.2021.103330>
- [121] Bai L, Lu S, Zhao Q, Chen L, Jiang Y, Jia C, Chen S. Low-energy-consuming CO₂ capture by liquid-liquid biphasic absorbents of EMEA/DEEA/PX. *Chem Eng J*. 2022;450:138490. <https://doi.org/10.1016/j.cej.2022.138490>

- [122] Yu Y, Shen Y, Zhou X, Liu F, Zhang S, Lu S, Ye J, Li S, Chen J, Li W. Relationship between tertiary amine's physical property and biphasic solvent's CO₂ absorption performance: quantum calculation and experimental demonstration. *Chem Eng J.* 2022;428:131241. <https://doi.org/10.1016/j.cej.2021.131241>
- [123] Zhou X, Liu C, Fan Y, Zhang L, Tang S, Mo S, Zhu Y, Zhu Z. Energy-efficient carbon dioxide capture using a novel low-viscous secondary amine-based nonaqueous biphasic solvent: Performance, mechanism, and thermodynamics. *Energy.* 2022;255:124570. <https://doi.org/10.1016/j.energy.2022.124570>
- [124] Jiang C, Chen H, Wang J, Shen Y, Ye J, Zhang S, Wang L, Chen J. Phase splitting agent regulated biphasic solvent for efficient CO₂ capture with a low heat duty. *Environ Sci Technol* 2020;54(12):7601-10. <https://doi.org/10.1021/acs.est.9b07923>
- [125] Ciftja AF, Hartono A, Svendsen HF. Carbamate stability measurements in amine/CO₂/water systems with Nuclear Magnetic Resonance (NMR) spectroscopy. *Energy Procedia* 2014;63:633-9. <https://doi.org/10.1016/j.egypro.2014.11.068>
- [126] Zhang S, Shen Y, Shao P, Chen J, Wang L. Kinetics, thermodynamics, and mechanism of a novel biphasic solvent for CO₂ capture from flue gas. *Environ Sci Technol* 2018;52(6):3660-8. <https://doi.org/10.1021/acs.est.7b05936>
- [127] Versteeg GF, Blauwhoff PM, van Swaaij WP. The effect of diffusivity on gas-liquid mass transfer in stirred vessels. Experiments at atmospheric and elevated pressures. *Chem Eng Sci* 1987;42(5):1103-19. [https://doi.org/10.1016/0009-2509\(87\)80060-X](https://doi.org/10.1016/0009-2509(87)80060-X)
- [128] Nookuea W, Tan Y, Li H, Thorin E, Yan J. Impacts of thermo-physical properties of gas and liquid phases on design of absorber for CO₂ capture using monoethanolamine. *Int J Greenh Gas Control* 2016;52:190-200. <https://doi.org/10.1016/j.ijggc.2016.07.012>

- [129] Zhou X, Shen Y, Liu F, Ye J, Wang X, Zhao J, Zhang S, Wang L, Li S, Chen J. A novel dual-stage phase separation process for CO₂ absorption into a biphasic solvent with low energy penalty. *Environ Sci Technol* 2021;55(22):15313-22. <https://doi.org/10.1021/acs.est.1c01622>
- [130] Madejski P, Chmiel K, Subramanian N, Kuś T. Methods and techniques for CO₂ capture: Review of potential solutions and applications in modern energy technologies. *Energies* 2022;15(3):887. <https://doi.org/10.3390/en15030887>
- [131] U.S. Energy Information Administration, International Energy Outlook 2023 (IEO2023). <https://www.eia.gov/aeo; 2023> [accessed 01 October 2024]
- [132] Global CCS Institute. <https://www.globalccsinstitute.com/wp-content/uploads/2024/01/Global-Status-of-CCS-Report-1.pdf>; 2024 [accessed 01 October 2024]
- [133] Global CCS Institute. <https://www.globalccsinstitute.com/wp-content/uploads/2021/03/Global-Status-of-CCS-Report-English.pdf> [accessed 01 October 2024]

2.5 Objective of the research work

This thesis aims to select novel aqueous amine blends and study their performance for post-combustion CO₂ capture. The overall point-wise objectives of this thesis work are as follows:

- ❖ The prime objective of this thesis work is to evaluate CO₂ absorption-desorption performance for the selected novel aqueous amine blends of BAE+DMAE, HMDA+DMAE, and TETA+DMAE for post-combustion CO₂ capture.
- ❖ The first objective is validating the experimental setup and estimating equilibrium CO₂ loading during CO₂ absorption investigation.
- ❖ Analyzing the effect of process parameters on equilibrium CO₂ loading as per the chosen range of operating conditions.
- ❖ Developing an empirical model for the validation of CO₂ absorption results.
- ❖ To evaluate the cyclic capacity, cyclic equilibrium CO₂ loading, heat duty, and regeneration efficiency of the novel blends and their comparison with the benchmark MEA during the CO₂ desorption study.
- ❖ To examine pH, density, CO₂ absorption and desorption rates.
- ❖ Proposing reaction mechanism for novel aqueous amines of BAE+DMAE, HMDA+DMAE, and TETA+DMAE when reacting with CO₂ and their validation by ¹³C NMR and FTIR characterization.
- ❖ To study the heat of CO₂ absorption (ΔH_{abs}) for all novel aqueous amine blends.
- ❖ Modeling and optimization of equilibrium CO₂ loading by the RSM software.
- ❖ To study the toxic behavior of the chosen novel amine blends.

References

- [1] Gautam A, Mondal MK. Post-combustion capture of CO₂ using novel aqueous triethylenetetramine and 2-dimethylaminoethanol amine blend: equilibrium CO₂ loading-empirical model and optimization, CO₂ desorption, absorption heat, and ¹³C NMR analysis. *Fuel*. 2023;331:125864. <https://doi.org/10.1016/j.fuel.2022.125864>
- [2] Gautam A, Mondal MK. Novel aqueous amine blend of 2-(Butylamino) ethanol and 2-Dimethylaminoethanol for CO₂ capture: Equilibrium CO₂ loading, RSM optimization, desorption study, characterization and toxicity assessment. *Sep Purif Technol*. 2023;322:124279. <https://doi.org/10.1016/j.seppur.2023.124279>
- [3] Gautam A, Mondal MK. Review of recent trends and various techniques for CO₂ capture: Special emphasis on biphasic amine solvents. *Fuel*. 2023;334:126616. <https://doi.org/10.1016/j.fuel.2022.126616>
- [4] Gautam A, Mondal MK. Post-combustion CO₂ absorption-desorption performance of novel aqueous binary amine blend of Hexamethylenediamine (HMDA) and 2-Dimethylaminoethanol (DMAE). *Energy*. 2024;296:130982. <https://doi.org/10.1016/j.energy.2024.130982>
- [5] He X, He H, Barzagli F, Amer MW, Li CE, Zhang R. Analysis of the energy consumption in solvent regeneration processes using binary amine blends for CO₂ capture. *Energy*. 2023;270:126903. <https://doi.org/10.1016/j.energy.2023.126903>
- [6] Song C, Liu Q, Deng S, Li H, Kitamura Y. Cryogenic-based CO₂ capture technologies: State-of-the-art developments and current challenges. *Renew Sustain Energy Rev* 2019;101:265-78. <https://doi.org/10.1016/j.rser.2018.11.018>

- [7] Borhani TN, Wang M. Role of solvents in CO₂ capture processes: The review of selection and design methods. *Renew Sustain Energy Rev.* 2019;114:109299. <https://doi.org/10.1016/j.rser.2019.109299>
- [8] Ramdin M, de Loos TW, Vlucht TJ. State-of-the-art of CO₂ capture with ionic liquids. *Ind Eng Chem Res* 2012;51(24):8149-77. <https://doi.org/10.1021/ie3003705>
- [9] Kumar S, Mondal MK. Selection of efficient absorbent for CO₂ capture from gases containing low CO₂. *Korean J Chem Eng* 2020;37:231-9. <https://doi.org/10.1007/s11814-019-0440-6>
- [10] Kim YE, Yun SH, Choi JH, Nam SC, Park SY, Jeong SK, Yoon YI. Comparison of the CO₂ absorption characteristics of aqueous solutions of diamines: absorption capacity, specific heat capacity, and heat of absorption. *Energy Fuels* 2015;29:2582-90. <https://doi.org/10.1021/ef500561a>
- [11] Wai SK, Nwaoha C, Saiwan C, Idem R, Supap T. Absorption heat, solubility, absorption and desorption rates, cyclic capacity, heat duty, and absorption kinetic modeling of AMP–DETA blend for post–combustion CO₂ capture. *Sep Purif Technol* 2018;194:89-95. <https://doi.org/10.1016/j.seppur.2017.11.024>
- [12] Mondal BK, Bandyopadhyay SS, Samanta AN. Kinetics of CO₂ absorption in aqueous hexamethylenediamine. *Int J Greenh Gas Control* 2017;56:116-25. <https://doi.org/10.1016/j.ijggc.2016.11.023>
- [13] Vaidya PD, Kenig EY. CO₂-alkanolamine reaction kinetics: a review of recent studies. *Chem Eng Technol* 2007;30:1467-74. <https://doi.org/10.1002/ceat.200700268>
- [14] Zhao S, Wang Y, Zhu K, Zhao D, Song Q. Improving the absorption load, high viscosity, and regeneration efficiency of CO₂ capture using a novel tri-solvent biphasic solvents of TETA-AMP-1DMA2P. *Environ Sci Pollut Res* 2022;29:84903-15. <https://doi.org/10.1007/s11356-022-21822-6>

- [15] Lv B, Guo B, Zhou Z, Jing G. Mechanisms of CO₂ capture into monoethanolamine solution with different CO₂ loading during the absorption/desorption processes. *Environ Sci Technol* 2015;49:10728-35. <https://doi.org/10.1021/acs.est.5b02356>
- [16] Muchan P, Saiwan C, Narku-Tetteh J, Idem R, Supap T, Tontiwachwuthikul P. Screening tests of aqueous alkanolamine solutions based on primary, secondary, and tertiary structure for blended aqueous amine solution selection in post combustion CO₂ capture. *Chem Eng Sci* 2017;170:574-82. <https://doi.org/10.1016/j.ces.2017.02.031>
- [17] Hu X, Huang J, He X, Luo Q, Li CE, Zhou C, Zhang R. Analyzing the potential benefits of trio-amine systems for enhancing the CO₂ desorption processes. *Fuel*. 2022;316:123216. <https://doi.org/10.1016/j.fuel.2022.123216>
- [18] Li T, Yang C, Tantikhajorngosol P, Sema T, Shi H, Tontiwachwuthikul P. Experimental investigations and the modeling approach for CO₂ solubility in aqueous blended amine systems of monoethanolamine, 2-amino-2-methyl-1-propanol, and 2-(butylamino) ethanol. *Environ Sci Pollut Res* 2022;29:69402-23. <https://doi.org/10.1007/s11356-022-20411-x>
- [19] Sreedhar I, Nahar T, Venugopal A, Srinivas B. Carbon capture by absorption–Path covered and ahead. *Renew Sustain Energy Rev* 2017;76:1080-107. <https://doi.org/10.1016/j.rser.2017.03.109>
- [20] Chen M, Li M, Zhang F, Hu X, Wu Y. Fast and efficient CO₂ absorption in non-aqueous tertiary amines promoted by ethylene glycol. *Energy Fuels* 2022;36:4830-6. <https://doi.org/10.1021/acs.energyfuels.2c00215>
- [21] Kim YE, Moon SJ, Yoon YI, Jeong SK, Park KT, Bae ST, Nam SC. Heat of absorption and absorption capacity of CO₂ in aqueous solutions of amine containing multiple amino groups. *Sep Purif Technol* 2014;122:112-8. <https://doi.org/10.1016/j.seppur.2013.10.030>

- [22] Mazari SA, Kang TH, Devkota S, Cha JY, Shin BJ, Mun JH, Kim KM, Lee U, Moon JH. Investigating the effect of blending of diamine and alkanolamine for CO₂ capture: Experiment and thermodynamic modeling of CO₂-AEEA-DEA-H₂O system. *Chem Eng J.* 2023;470:144141. <https://doi.org/10.1016/j.cej.2023.144141>
- [23] Ji L, Yu H, Li K, Yu B, Grigore M, Yang Q, Wang X, Chen Z, Zeng M, Zhao S. Integrated absorption-mineralisation for low-energy CO₂ capture and sequestration. *Appl Energy* 2018;225: 356-66. <https://doi.org/10.1016/j.apenergy.2018.04.108>
- [24] Cao F, Gao H, Xiong Q, Liang Z. Experimental studies on mass transfer performance for CO₂ absorption into aqueous N, N-dimethylethanolamine (DMEA) based solutions in a PTFE hollow fiber membrane contactor. *Int J Greenh Gas Control* 2019;82:210-7. <https://doi.org/10.1016/j.ijggc.2018.12.011>
- [25] Singh S, Pandey D, Mondal MK. New Experimental Data on Equilibrium CO₂ Loading into Aqueous 3-Dimethyl Amino-1-propanol and 1, 5-Diamino-2-methylpentane Blend: Empirical Model and CO₂ Absorption Enthalpy. *J Chem Eng Data* 2020;66:740-8. <https://doi.org/10.1021/acs.jced.0c00851>
- [26] Zhang J, Qiao Y, Wang W, Misch R, Hussain K, Agar DW. Development of an energy-efficient CO₂ capture process using thermomorphic biphasic solvents. *Energy Procedia* 2013;37:1254-61. <https://doi.org/10.1016/j.egypro.2013.05.224>
- [27] Xiao M, Cui D, Liu H, Tontiwachwuthikul P, Liang Z. A new model for correlation and prediction of equilibrium CO₂ solubility in N-methyl-4-piperidinol solvent. *AIChE J* 2017;63:3395-403. <https://doi.org/10.1002/aic.15709>
- [28] Nwaoha C, Saiwan C, Supap T, Idem R, Tontiwachwuthikul P, Rongwong W, Al-Marri MJ, Benamor A. Carbon dioxide (CO₂) capture performance of aqueous tri-solvent blends containing 2-amino-2-methyl-1-propanol (AMP) and methyldiethanolamine

- (MDEA) promoted by diethylenetriamine (DETA). *Int J Greenh Gas Control* 2016;53:292-04. <https://doi.org/10.1016/j.ijggc.2016.08.012>
- [29] El Hadri N, Quang DV, Goetheer EL, Zahra MR. Aqueous amine solution characterization for post-combustion CO₂ capture process. *Appl Energy* 2017;185:1433-49. <https://doi.org/10.1016/j.apenergy.2016.03.043>
- [30] Chowdhury FA, Okabe H, Shimizu S, Onoda M, Fujioka Y. Development of novel tertiary amine absorbents for CO₂ capture. *Energy Procedia* 2009;1:1241-8. <https://doi.org/10.1016/j.egypro.2009.01.163>
- [31] Ling H, Gao H, Liang Z. Comprehensive solubility of N₂O and mass transfer studies on an effective reactive N, N-dimethylethanolamine (DMEA) solvent for post-combustion CO₂ capture. *Chem Eng J* 2019;355:369-79. <https://doi.org/10.1016/j.cej.2018.08.147>
- [32] Sharif M, Zhang T, Wu X, Yu Y, Zhang Z. Evaluation of CO₂ absorption performance by molecular dynamic simulation for mixed secondary and tertiary amines. *Int J Greenh Gas Control*. 2020;97:103059. <https://doi.org/10.1016/j.ijggc.2020.103059>
- [33] Conway W, Bruggink S, Beyad Y, Luo W, Melián-Cabrera I, Puxty G, Feron P. CO₂ absorption into aqueous amine blended solutions containing monoethanolamine (MEA), N, N-dimethylethanolamine (DMEA), N, N-diethylethanolamine (DEEA) and 2-amino-2-methyl-1-propanol (AMP) for post-combustion capture processes. *Chem Eng Sci* 2015;126:446-54. <https://doi.org/10.1016/j.ces.2014.12.053>
- [34] Liu S, Ling H, Gao H, Tontiwachwuthikul P, Liang Z. Better choice of tertiary alkanolamines for postcombustion CO₂ capture: structure with linear alkanol chain instead of branched. *Ind Eng Chem Res* 2019;58(33):15344-52. <https://doi.org/10.1021/acs.iecr.9b02244>

- [35] Chowdhury FA, Yamada H, Higashii T, Goto K, Onoda M. CO₂ capture by tertiary amine absorbents: a performance comparison study. *Ind Eng Chem Res* 2013;52(24):8323-31. <https://doi.org/10.1021/ie400825u>
- [36] Xiao M, Liu H, Idem R, Tontiwachwuthikul P, Liang Z. A study of structure–activity relationships of commercial tertiary amines for post-combustion CO₂ capture. *Appl Energy* 2016;184:219-29. <https://doi.org/10.1016/j.apenergy.2016.10.006>
- [37] Gao H, Wu Z, Liu H, Luo X, Liang Z. Experimental studies on the effect of tertiary amine promoters in aqueous monoethanolamine (MEA) solutions on the absorption/stripping performances in post-combustion CO₂ capture. *Energy Fuels* 2017;31(12):13883-91. <https://doi.org/10.1021/acs.energyfuels.7b02390>
- [38] Gao G, Xu B, Gao X, Jiang W, Zhao Z, Li X, Luo C, Wu F, Zhang L. New insights into the structure-activity relationship for CO₂ capture by tertiary amines from the experimental and quantum chemical calculation perspectives. *Chem Eng J*. 2023;473:145277. <https://doi.org/10.1016/j.cej.2023.145277>
- [39] Baltar A, Gómez-Díaz D, Navaza JM, Rumbo A. Absorption and regeneration studies of chemical solvents based on dimethylethanolamine and diethylethanolamine for carbon dioxide capture. *AIChE J*. 2019;66(1):e16770. <https://doi.org/10.1002/aic.16770>
- [40] Shen Y, Jiang C, Zhang S, Chen J, Wang L, Chen J. Biphasic solvent for CO₂ capture: Amine property-performance and heat duty relationship. *Appl Energy* 2018;230:726-33. <https://doi.org/10.1016/j.apenergy.2018.09.005>
- [41] Singto S, Supap T, Idem R, Tontiwachwuthikul P, Tantayanon S, Al-Marri MJ, Benamor A. Synthesis of new amines for enhanced carbon dioxide (CO₂) capture performance: The effect of chemical structure on equilibrium solubility, cyclic capacity, kinetics of absorption and regeneration, and heats of absorption and regeneration. *Sep Purif Technol* 2016;167:97-107. <https://doi.org/10.1016/j.seppur.2016.05.002>

- [42] Narku-Tetteh J, Muchan P, Saiwan C, Supap T, Idem R. Selection of components for formulation of amine blends for post combustion CO₂ capture based on the side chain structure of primary, secondary and tertiary amines. *Chem Eng Sci* 2017;170:542-60. <https://doi.org/10.1016/j.ces.2017.02.036>
- [43] Xiao M, Liu H, Gao H, Liang Z. CO₂ absorption with aqueous tertiary amine solutions: Equilibrium solubility and thermodynamic modeling. *J Chem Thermodyn* 2018;122:170-82. <https://doi.org/10.1016/j.jct.2018.03.020>
- [44] Mokhtari Z, Pakravesht A, Zarei H. High-pressure densities of 2-(dimethylamino) ethanol and 2-(diethylamino) ethanol: Measurement and modeling with new modified Tait and PC-SAFT equations of state. *Fluid Ph Equilibria*. 2023;572:113825. <https://doi.org/10.1016/j.fluid.2023.113825>
- [45] Sharif M, Fan H, Wu X, Yu Y, Zhang T, Zhang Z. Assessment of novel solvent system for CO₂ capture applications. *Fuel*. 2023;337:127218. <https://doi.org/10.1016/j.fuel.2022.127218>
- [46] Zhang P, Xu R, Li H, Gao H, Liang Z. Mass transfer performance for CO₂ absorption into aqueous blended DMEA/MEA solution with optimized molar ratio in a hollow fiber membrane contactor. *Sep Purif Technol* 2019;211:628-36. <https://doi.org/10.1016/j.seppur.2018.10.034>
- [47] Ling H, Liu S, Wang T, Gao H, Liang Z. Characterization and correlations of CO₂ absorption performance into aqueous amine blended solution of monoethanolamine (MEA) and N, N-dimethylethanolamine (DMEA) in a packed column. *Energy Fuels* 2019;33:7614-25. <https://doi.org/10.1021/acs.energyfuels.9b01764>
- [48] Jiang W, Luo X, Gao H, Liang Z, Liu B, Tontiwachwuthikul P, Hu X. A comparative kinetics study of CO₂ absorption into aqueous DEEA/MEA and DMEA/MEA blended solutions. *AIChE J* 2018;64:1350-8. <https://doi.org/10.1002/aic.16024>

- [49] Buvik V, Vevelstad SJ, Brakstad OG, Knuutila HK. Stability of structurally varied aqueous amines for CO₂ capture. *Ind Eng Chem Res* 2021;60:5627-38. <https://doi.org/10.1021/acs.iecr.1c00502>
- [50] Brúder P, Lauritsen KG, Mejdell T, Svendsen HF. CO₂ capture into aqueous solutions of 3-methylaminopropylamine activated dimethyl-monoethanolamine. *Chem Eng Sci* 2012;75:28-37. <https://doi.org/10.1016/j.ces.2012.03.005>
- [51] Zhang J, Tong D, Fennell PS, Trusler JM. Solubility of CO₂ in aqueous amine solutions: A study to select solvents for carbon capture from natural-gas power plant. In *Proceedings of the 4th International Gas Processing Symposium* 2015;1-10. <https://doi.org/10.1016/B978-0-444-63461-0.50001-8>
- [52] Concepción EI, Moreau A, Segovia JJ, Pérez Y, Arroyave JD, Martín MC. A comparative study of thermophysical properties of amine aqueous solutions for CO₂ mitigation. <https://doi.org/10.52202/069564-0009>
- [53] Delavari M, Khajenoori M, Zoghi AT. Equilibrium absorption of CO₂ in aqueous solution of N-dimethylamino ethanol and 2-(ethylamino) ethanol, measuring and thermodynamic modeling. *J Chem Thermodyn.* 2023;186:107142. <https://doi.org/10.1016/j.jct.2023.107142>
- [54] Kim J, Kim K, Lim H, Kang JH, Park HS, Park J, Song H. Structural investigation of aqueous amine solutions for CO₂ capture: CO₂ loading, cyclic capacity, absorption–desorption rate, and pKa. *J Environ Chem Eng.* 2024;12(3):112664. <https://doi.org/10.1016/j.jece.2024.112664>
- [55] Rahimi A, Zoghi AT, Feyzi F, Jalili AH. Experimental study of density, viscosity and equilibrium carbon dioxide solubility in some aqueous alkanolamine solutions. *J Solution Chem.* 2019;48:489-501. <https://doi.org/10.1007/s10953-019-00872>

References

- [1] Gautam A, Kumar Mondal M, 2023. Post-combustion capture of CO₂ using novel aqueous Triethylenetetramine and 2-Dimethylaminoethanol amine blend: Equilibrium CO₂ loading-empirical model and optimization, CO₂ desorption, absorption heat, and ¹³C NMR analysis. *Fuel*. 331, 125864. <https://doi.org/10.1016/j.fuel.2022.125864>.
- [2] Gautam A, Mondal MK, 2023. Review of recent trends and various techniques for CO₂ capture: Special emphasis on biphasic amine solvents. *Fuel*. 334, 126616. <https://doi.org/10.1016/j.fuel.2022.126616>.
- [3] Agnihotri N, Gupta GK, Mondal MK. Thermo-kinetic analysis, thermodynamic parameters and comprehensive pyrolysis index of Melia azedarach sawdust as a genesis of bioenergy. *Biomass Conv Bioref* 2024;14:1863–80. <https://doi.org/10.1007/s13399-022-02524-y>.
- [4] Wang R, Zhao H, Wang Y, Qi C, Zhang S, Wang L, Li M, 2022. Development of biphasic solvent for CO₂ capture by tailoring the polarity of amine solution. *Fuel*. 325, 124885. <https://doi.org/10.1016/j.fuel.2022.124885>.
- [5] Ahmed RE, Wiheeb AD. Enhancement of carbon dioxide absorption into aqueous potassium carbonate by adding amino acid salts. *Mater Today Proc* 20:2020;611-6. <https://doi.org/10.1016/j.matpr.2019.09.198>.
- [6] Agnihotri N, Mondal MK, 2023. Comparison of non-catalytic and in-situ catalytic pyrolysis of Melia azedarach sawdust. *J Anal Appl Pyrolysis* 172, 106006. <https://doi.org/10.1016/j.jaap.2023.106006>.

- [7] Kim J, Lee J, Lee Y, Kim H, Kim E, Lee KS, 2019. Evaluation of aqueous polyamines as CO₂ capture solvents. *Energy*. 187, 115908. <https://doi.org/10.1016/j.energy.2019.115908>.
- [8] Pandey D, Kumar Mondal M, 2021. Thermodynamic modeling and new experimental CO₂ solubility into aqueous EAE and AEEA blend, heat of absorption, cyclic absorption capacity and desorption study for post-combustion CO₂ capture. *Chem Eng J* 410, 128334. <https://doi.org/10.1016/j.cej.2020.128334>.
- [9] Xiao M, Liu H, Gao H, Liang Z. Liang. CO₂ absorption with aqueous tertiary amine solutions: Equilibrium solubility and thermodynamic modeling. *J Chem Thermodyn* 2018;122:170–82. <https://doi.org/10.1016/j.jct.2018.03.020>.
- [10] Chowdhury FA, Okabe H, Shimizu S, Onoda M, Fujioka Y. Development of novel tertiary amine absorbents for CO₂ capture. *Energy Procedia* 2009;1:1241–8. <https://doi.org/10.1016/j.egypro.2009.01.163>.
- [11] Kim YE, Park JH, Yun SH, Nam SC, Jeong SK, Yoon Yi. Carbon dioxide absorption using a phase transitional alkanolamine – alcohol mixture. *J Ind Eng Chem* 2014;20:1486–92. <https://doi.org/10.1016/j.jiec.2013.07.036>.
- [12] Lv B, Guo B, Zhou Z, Jing G. Mechanisms of CO₂ capture into monoethanolamine solution with different CO₂ loading during the absorption/desorption processes. *Environ Sci Technol* 2015;49:10728–35. <https://doi.org/10.1021/acs.est.5b02356>.
- [13] Muchan P, Saiwan C, Narku-Tetteh J, Idem R, Supap T, Tontiwachwuthikul P. Screening tests of aqueous alkanolamine solutions based on primary, secondary, and tertiary structure for blended aqueous amine solution selection in post combustion CO₂ capture. *Chem Eng Sci* 2017;170:574–82. <https://doi.org/10.1016/j.ces.2017.02.031>.

- [14] Huang Q, Jing G, Zhou X, Lv B, Zhou Z. A novel biphasic solvent of amino-functionalized ionic liquid for CO₂ capture: High efficiency and regenerability. *J CO₂ Util* 2018;25:22–30. <https://doi.org/10.1016/j.jcou.2018.03.001>.
- [15] Ji L, Yu H, Li K, Yu B, Grigore M, Yang Q, et al. Integrated absorption-mineralisation for low-energy CO₂ capture and sequestration. *Appl Energy* 2018;225:356–66. <https://doi.org/10.1016/j.apenergy.2018.04.108>.
- [16] Chen M, Li M, Zhang F, Hu X, Wu Y. Fast and Efficient CO₂ Absorption in Non-aqueous Tertiary Amines Promoted by Ethylene Glycol. *Energy Fuels* 2022;36:4830–6. <https://doi.org/10.1021/acs.energyfuels.2c00215>.
- [17] Gao G, Li X, Jiang W, Zhao Z, Xu Y, Wu F, Luo C, Zhang L, 2022. Improved quasi-cycle capacity method based on microcalorimetry strategy for the fast screening of amino acid salt absorbents for CO₂ capture. *Sep Purif Technol* 289, 120767. <https://doi.org/10.1016/j.seppur.2022.120767>.
- [18] Zhou X, Liu F, Lv B, Zhou Z, Jing G. Evaluation of the novel biphasic solvents for CO₂ capture: Performance and mechanism. *Int J Greenh Gas Control* 2017;60:120–8. <https://doi.org/10.1016/j.ijggc.2017.03.013>.
- [19] Agnihotri N, Mondal MK, 2023. Thermal analysis, kinetic behavior, reaction modeling, and comprehensive pyrolysis index of soybean stalk pyrolysis. *Biomass Conv Bioref*. <https://doi.org/10.1007/s13399-023-03807-8>.
- [20] Kumar S, Padhan R, Mondal MK. Equilibrium solubility of CO₂ in aqueous blend of 2-(diethylamine) ethanol and 2-(2-aminoethylamine) ethanol. *J Chem Eng Data* 2018;63:1163-9. <https://doi.org/10.1021/acs.jced.7b00544>.

- [21] Kumar S, Mondal MK. Equilibrium solubility measurement and modeling of CO₂ absorption in aqueous blend of 2-(diethyl amino) ethanol and ethylenediamine. *J Chem Eng Data* 2020;65:523-31. <https://doi.org/10.1021/acs.jced.9b00699>.
- [22] Kumar S, Mondal MK. Selection of efficient absorbent for CO₂ capture from gases containing low CO₂. *Korean J Chem Eng* 2020;37:231–9. <https://doi.org/10.1007/s11814-019-0440-6>
- [23] Perumal M, Jayaraman D, Balraj A. Selection of efficient absorbent for CO₂ capture from gases containing low CO₂. *Korean J Chem Eng* 2020;37:231-9. <https://doi.org/10.1007/s11814-019-0440-6>.
- [24] Li T, Yang C, Tantikhajongosol P, Sema T, Shi H, Tontiwachwuthikul P, 2021. Experimental studies on CO₂ absorption and solvent recovery in aqueous blends of monoethanolamine and tetrabutylammonium hydroxide. *Chemosphere*. 276, 130159. <https://doi.org/10.1016/j.chemosphere.2021.130159>.
- [25] Pandey D, Mondal MK, 2020. Equilibrium CO₂ solubility in the aqueous mixture of MAE and AEEA: Experimental study and development of modified thermodynamic model. *Fluid Phase Equilib*. 522, 112766. <https://doi.org/10.1016/j.fluid.2020.112766>.
- [26] Singh S, Pandey D, Mondal MK. New Experimental Data on Equilibrium CO₂ Loading into Aqueous 3-Dimethyl Amino-1-propanol and 1,5-Diamino-2-methylpentane Blend: Empirical Model and CO₂ Absorption Enthalpy. *J Chem Eng Data* 2021;66:740–8. <https://doi.org/10.1021/acs.jced.0c00851>.
- [27] Zhang R, He X, Liu T, Li C, Xiao M, Ling H, Hu X, Zhang X, Tang F, Luo HA, 2022. Thermodynamic studies for improving the prediction of CO₂ equilibrium solubility in

aqueous 2-dimethylamino-2-methyl-1-propanol. *Sep Purif Technol* 295, 121292. <https://doi.org/10.1016/j.seppur.2022.121292>.

[28] Machida H, Oba K, Tomikawa T, Esaki T, Yamaguchi T, Horizoe H. Development of phase separation solvent for CO₂ capture by aqueous (amine + ether) solution. *J Chem Thermodyn* 2017;113:64–70. <https://doi.org/10.1016/j.jct.2017.05.043>.

[29] Ping T, Dong Y, Shen S. Shen, Energy-efficient CO₂ capture using nonaqueous absorbents of secondary alkanolamines with a 2-butoxyethanol cosolvent. *ACS Sustain Chem Eng* 2020;8:18071–82. <https://doi.org/10.1021/acssuschemeng.0c06345>.

[30] Dong Y, Ping T, Shen S. Solubility of CO₂ in nonaqueous system of 2-(butylamino) ethanol with 2-butoxyethanol: Experimental data and model representation. *Chinese J Chem Eng* 2022;41:441–8. <https://doi.org/10.1016/j.cjche.2021.11.003>.

[31] Ping T, Dong Y, Shen S, 2020. Densities, viscosities and spectroscopic study of partially CO₂-loaded nonaqueous blends of 2-butoxyethanol with 2-(ethylamino) ethanol and 2-(butylamino) ethanol at temperatures of (293.15 to 353.15) K. *J Mol Liq* 312, 113389. <https://doi.org/10.1016/j.molliq.2020.113389>.

[32] Liang ZH, Rongwong W, Liu H, Fu K, Gao H, Cao F, Zhang R, Sema T, Henni A, Sumon K, Nath D. Recent progress and new developments in post-combustion carbon-capture technology with amine based solvents. *Int J Greenh Gas Control* 2015;40:26-54. <https://doi.org/10.1016/j.ijggc.2015.06.017>.

[33] El Hadri N, Quang DV, Goetheer ELV, Abu Zahra MRM. Aqueous amine solution characterization for post-combustion CO₂ capture process. *Appl Energy* 2017;185:1433-49. <https://doi.org/10.1016/j.apenergy.2016.03.043>.

- [34] Wai SK, Saiwan C, Idem R, Supap T, Nwaoha C. Carbon Dioxide (CO₂) Solubility in Diethylenetriamine and 2-Amino-2-Methyl-1-Propanal (DETA-AMP) Solvent System for Amine-Based CO₂ Capture in Flue Gas from Coal Combustion. *Energy Procedia* 2017;114:1973–9. <https://doi.org/10.1016/j.egypro.2017.03.1329>.
- [35] Liu H, Li M, Idem R, (PT) Tontiwachwuthikul P, Liang Z. Analysis of solubility, absorption heat and kinetics of CO₂ absorption into 1-(2-hydroxyethyl)pyrrolidine solvent. *Chem Eng Sci* 2017;162:120–30. <https://doi.org/10.1016/j.ces.2016.12.070>.
- [36] Aronu UE, Gondal S, Hessen ET, Haug-Warberg T, Hartono A, Hoff KA, et al. Solubility of CO₂ in 15, 30, 45 and 60 mass% MEA from 40 to 120 °C and model representation using the extended UNIQUAC framework. *Chem Eng Sci* 2011;66:6393–406. <https://doi.org/10.1016/j.ces.2011.08.042>.
- [37] Barzagli F, Mani F, Peruzzini M. A Comparative Study of the CO₂ Absorption in Some Solvent-Free Alkanolamines and in Aqueous Monoethanolamine (MEA). *Environ Sci Technol* 2016;50:7239–46. <https://doi.org/10.1021/acs.est.6b00150>.
- [38] Afari DB, Coker J, Narku-Tetteh J, Idem R. Comparative Kinetic Studies of Solid Absorber Catalyst (K/MgO) and Solid Desorber Catalyst (HZSM-5)-Aided CO₂ Absorption and Desorption from Aqueous Solutions of MEA and Blended Solutions of BEA-AMP and MEA-MDEA. *Ind Eng Chem Res* 2018;57:15824–39. <https://doi.org/10.1021/acs.iecr.8b02931>.
- [39] Srisang W, Pouryousefi F, Osei PA, Decardi-Nelson B, Akachuku A, Tontiwachwuthikul P, Idem R. Evaluation of the heat duty of catalyst-aided amine-based post combustion CO₂ capture. *Chem Eng Sci* 2017;170:48–57. <https://doi.org/10.1016/j.ces.2017.01.049>.

- [40] Cao F, Gao H, Xiong Q, Liang Z. Experimental studies on mass transfer performance for CO₂ absorption into aqueous N,N-dimethylethanolamine (DMEA) based solutions in a PTFE hollow fiber membrane contactor. *Int J Greenh Gas Control* 2019;82:210–7. <https://doi.org/10.1016/j.ijggc.2018.12.011>.
- [41] Ling H, Gao H, Liang Z. Comprehensive solubility of N₂O and mass transfer studies on an effective reactive N,N-dimethylethanolamine (DMEA) solvent for post-combustion CO₂ capture. *Chem Eng J* 2019;355:369–79. <https://doi.org/10.1016/j.cej.2018.08.147>.
- [42] Nwaoha C, Saiwan C, Supap T, Idem R, Tontiwachwuthikul P, Rongwong W, et al. Carbon dioxide (CO₂) capture performance of aqueous tri-solvent blends containing 2-amino-2-methyl-1-propanol (AMP) and methyldiethanolamine (MDEA) promoted by diethylenetriamine (DETA). *Int J Greenh Gas Control* 2016;53:292–304. <https://doi.org/10.1016/j.ijggc.2016.08.012>.
- [43] Hwang SJ, Kim J, Kim H, Lee KS. Solubility of Carbon Dioxide in Aqueous Solutions of Three Secondary Amines: 2-(Butylamino)ethanol, 2-(Isopropylamino)ethanol, and 2-(Ethylamino)ethanol Secondary Alkanolamine Solutions. *J Chem Eng Data* 2017;62:2428–35. <https://doi.org/10.1021/acs.jced.7b00364>.
- [44] Barzagli F, Giorgi C, Mani F, Peruzzini M. Reversible carbon dioxide capture by aqueous and non-aqueous amine-based absorbents: A comparative analysis carried out by ¹³C NMR spectroscopy. *Appl Energy* 2018;220:208–19. <https://doi.org/10.1016/j.apenergy.2018.03.076>.
- [45] Xiao M, Cui D, Liu H, Tontiwachwuthikul P, Liang Z. A new model for correlation and prediction of equilibrium CO₂ solubility in N-methyl-4-piperidinol solvent. *AIChE J* 2017;63:3395–403. <https://doi.org/10.1002/aic.15709>.

- [46] Zheng W, Yan Z, Zhang R, Jiang W, Luo X, Liang Z, Yang Q, Yu H, 2022. A study of kinetics, equilibrium solubility, speciation and thermodynamics of CO₂ absorption into benzylamine (BZA) solution. *Chem Eng Sci.* 251, 117452. <https://doi.org/10.1016/j.ces.2022.117452>.
- [47] Narku-Tetteh J, Muchan P, Saiwan C, Supap T, Idem R. Selection of components for formulation of amine blends for post combustion CO₂ capture based on the side chain structure of primary, secondary and tertiary amines. *Chem Eng Sci* 2017;170:542–60. <https://doi.org/10.1016/j.ces.2017.02.036>.
- [48] Ramezani R, Mazinani S, Di Felice R. Density, Viscosity, pH, Heat of Absorption, and CO₂ Loading Capacity of Methyldiethanolamine and Potassium Lysinate Blend Solutions. *J Chem Eng Data* 2021;66:1611–29. <https://doi.org/10.1021/acs.jced.0c00855>.
- [49] Gadaleta D, Vuković K, Toma C, Lavado GJ, Karmaus AL, Mansouri K, Kleinstreuer NC, Benfenati E, Roncaglioni A. SAR and QSAR modeling of a large collection of LD₅₀ rat acute oral toxicity data. *J Cheminform* 2019;11:1-6. <https://doi.org/10.1186/s13321-019-0383-2>.
- [50] Asgarifard P, Rahimi M, Tafreshi N. Response surface modelling of CO₂ capture by ammonia aqueous solution in a microchannel. *Can J Chem Eng* 2021;99:601–12. <https://doi.org/10.1002/cjce.23881>.
- [51] Gil MV, Martínez M, García S, Rubiera F, Pis JJ, Pevida C. Response surface methodology as an efficient tool for optimizing carbon adsorbents for CO₂ capture. *Fuel Process Technol* 2013;106:55–61. <https://doi.org/10.1016/j.fuproc.2012.06.018>.

- [52] Tang Q, Lau Y Bin, Hu S, Yan W, Yang Y, Chen T. Response surface methodology using Gaussian processes: Towards optimizing the trans-stilbene epoxidation over Co^{2+} -NaX catalysts. *Chem Eng J* 2010;156:423–31. <https://doi.org/10.1016/j.cej.2009.11.002>.
- [53] Khajeh M, Ghaemi A. Nanoclay montmorillonite as an adsorbent for CO_2 capture: Experimental and modeling. *J Chinese Chem Soc* 2020;67:253–66. <https://doi.org/10.1002/jccs.201900150>.
- [54] Amiri M, Shahhosseini S, Ghaemi A. Optimization of CO_2 Capture Process from Simulated Flue Gas by Dry Regenerable Alkali Metal Carbonate Based Adsorbent Using Response Surface Methodology. *Energy Fuels* 2017;31:5286–96. <https://doi.org/10.1021/acs.energyfuels.6b03303>.
- [55] Pashaei H, Ghaemi A, Nasiri M, Karami B. Experimental Modeling and Optimization of CO_2 Absorption into Piperazine Solutions Using RSM-CCD Methodology. *ACS Omega* 2020;5:8432–48. <https://doi.org/10.1021/acsomega.9b03363>.
- [56] Mourad AAHI, Mohammad AF, Al-Marzouqi AH, El-Naas MH, Al-Marzouqi MH, Altarawneh M, 2022. CO_2 capture and ions removal through reaction with potassium hydroxide in desalination reject brine: Statistical optimization. *Chem Eng Process: Process Intensif.* 170, 108722. <https://doi.org/10.1016/j.cep.2021.108722>.
- [57] Saeidi M, Ghaemi A, Tahvildari K, Derakhshi P. Exploiting response surface methodology (RSM) as a novel approach for the optimization of carbon dioxide adsorption by dry sodium hydroxide. *J Chinese Chem Soc* 2018;65:1465–75. <https://doi.org/10.1002/jccs.201800012>.
- [58] Das D, Meikap BC. Optimization of process condition for the preparation of amine-impregnated activated carbon developed for CO_2 capture and applied to methylene blue

adsorption by response surface methodology. *J Environ Sci Health* 2017;52:1164-72.
<https://doi.org/10.1080/10934529.2017.1356204>.

[59] Song C, Kitamura Y, Li S. Optimization of a novel cryogenic CO₂ capture process by response surface methodology (RSM). *J Taiwan Inst Chem Eng* 2014;45:1666–76.
<https://doi.org/10.1016/j.jtice.2013.12.009>.

[60] Sahraie S, Rashidi H, Valeh-e-Sheyda P. An optimization framework to investigate the CO₂ capture performance by MEA: Experimental and statistical studies using Box-Behnken design. *Process Saf Environ Prot* 2019;122:161-8.
<https://doi.org/10.1016/j.psep.2018.11.026>.

[61] Nuchitprasittichai A, Cremaschi S. Optimization of CO₂ capture process with aqueous amines using response surface methodology. *Comput Chem Eng* 2011;35:1521–31.
<https://doi.org/10.1016/j.compchemeng.2011.03.016>.

[62] Nuchitprasittichai A, Cremaschi S. Optimization of CO₂ Capture Process with Aqueous Amines - A Comparison of Two Simulation – Optimization Approaches. *Ind Eng Chem Res* 2013;52:10236-43. <https://doi.org/10.1021/ie3029366>.

[63] Rho SW, Yoo KP, Lee JS, Nam SC, Son JE, Min BM. Solubility of CO₂ in aqueous methyldiethanolamine solutions. *J Chem Eng Data* 1997;42:1161–4.
<https://doi.org/10.1021/je970097d>.

[64] Kim I, Svendsen HF. Heat of absorption of carbon dioxide (CO₂) in monoethanolamine (MEA) and 2-(aminoethyl)ethanolamine (AEEA) solutions. *Ind Eng Chem Res* 2007;46:5803–9. <https://doi.org/10.1021/ie0616489>.

- [65] Mathias PM, O'Connell JP. The Gibbs-Helmholtz equation and the thermodynamic consistency of chemical absorption data. *Ind Eng Chem Res* 2012;51:5090–7. <https://doi.org/10.1021/ie202668k>.
- [66] Caplow M. Kinetics of Carbamate Formation and Breakdown. *J Am Chem Soc* 1968;90:6795–803. <https://doi.org/10.1021/ja01026a041>.
- [67] Danckwerts PV. The reaction of CO₂ with ethanolamines. *Chem Eng Sci* 1979;34:443–6. [https://doi.org/10.1016/0009-2509\(79\)85087-3](https://doi.org/10.1016/0009-2509(79)85087-3).
- [68] Donaldson TL, Nguyen YN. Carbon Dioxide Reaction Kinetics and Transport in Aqueous Amine Membranes. *Ind Eng Chem Fundam* 1980;19:260–6. <https://doi.org/10.1021/i160075a005>.
- [69] Perinu C, Arstad B, Jens KJ. NMR spectroscopy applied to amine–CO₂–H₂O systems relevant for post-combustion CO₂ capture: A review. *Int J Greenh Gas Control* 2014;20:230–43. <https://doi.org/10.1016/j.ijggc.2013.10.029>.
- [70] Lee JI, Otto FD, Mather AE. Equilibrium Between Carbon Dioxide and Aqueous Monoethanolamine Solutions. *J Appl Chem Biotechnol* 1976;26:541–9. <https://doi.org/10.1002/jctb.5020260177>.
- [71] Shen KP, Li MH. Solubility of carbon dioxide in aqueous mixtures of monoethanolamine with methyldiethanolamine. *J Chem Eng Data* 1992;37:96–100. <https://doi.org/10.1021/je00005a025>.
- [72] Tong D, Trusler JM, Maitland GC, Gibbins J, Fennell PS. Solubility of carbon dioxide in aqueous solution of monoethanolamine or 2-amino-2-methyl-1-propanol: Experimental measurements and modeling. *Int J Greenh Gas Control* 2012;6:37–47. <https://doi.org/10.1016/j.ijggc.2011.11.005>.

- [73] Gao H, Wu Z, Liu H, Luo X, Liang Z. Experimental Studies on the Effect of Tertiary Amine Promoters in Aqueous Monoethanolamine (MEA) Solutions on the Absorption/Stripping Performances in Post-combustion CO₂ Capture. *Energy Fuels* 2017;31:13883-91. <https://doi.org/10.1021/acs.energyfuels.7b02390>.
- [74] Im J, Hong SY, Cheon Y, Lee J, Lee JS, Kim HS, et al. Steric hindrance-induced zwitterionic carbonates from alkanolamines and CO₂: Highly efficient CO₂ absorbents. *Energy Environ Sci* 2011;4:4284–9. <https://doi.org/10.1039/C1EE01801A>.
- [75] Agnihotri N, Mondal MK, 2023. Process parameter variation of Melia azedarach sawdust pyrolysis for fuel properties, physicochemical characterization, and in-depth speciation analysis. *Biomass Conv Bioref*. <https://doi.org/10.1007/s13399-023-04305-7>.
- [76] Wang X, Zheng K, Peng Z, Liu B, Jia X, Tian J, 2022. Exploiting proton masking to protect amino achieve efficient capture CO₂ by amino-acids deep eutectic solvents. *Sep Purif Technol* 299, 121787. <https://doi.org/10.1016/j.seppur.2022.121787>.
- [77] Ohno K, Matsumoto H, Yoshida H, Matsuura H, Iwaki T, Suda T. Vibrational Spectroscopic and ab Initio Studies on Conformations of the Chemical Species in a Reaction of Aqueous 2-(N, N-Dimethylamino) ethanol Solutions with Carbon Dioxide. Importance of Strong NH⁺...O Hydrogen Bonding. *J Phys Chem A*. 1998;102:8056–62. <https://doi.org/10.1021/jp982562z>.
- [78] Kortunov PV, Siskin M, Baugh LS, Calabro DC. In Situ Nuclear Magnetic Resonance Mechanistic Studies of Carbon Dioxide Reactions with Liquid Amines in Aqueous Systems: New Insights on Carbon Capture Reaction Pathways. *Energy Fuels* 2015;29:5919–39. <https://doi.org/10.1021/acs.energyfuels.5b00850>.

[79]

<https://www.fishersci.com/store/msds?partNumber=AC126720010&productDescription=N=METHYLDIETHANOLAMINE%2C+1KG&vendorId=VN00032119&countryCode=US&language=en> (accessed 02 February 2023).

[80] Spectrum. https://www.spectrumchemical.com/media/sds/AM156_AGHS.pdf (accessed 02 February 2023).

[81] Ted Pella. https://www.tedpella.com/SDS_html/18315_sds.pdf (accessed 02 February 2023).

[82] Oxford Lab Fine. [https://www.oxfordlabchem.com/msds/PIPERAZINE%20\(Anhydrous\).pdf](https://www.oxfordlabchem.com/msds/PIPERAZINE%20(Anhydrous).pdf) (accessed 02 February).

[83] Alliance Chemicals. <https://www.alliancechemicals.com/wp-content/uploads/2011/10/DEA-sds.pdf> (accessed 02 February).

[84] North Metal and Chemical. <https://northmetal.net/wp-content/uploads/Diethylaminoethanol-Diethylethanolamine-DEAE-DEEA-Pennad-150-C6H15NO-100-37-8-SDS.pdf> (accessed 02 February).

[85] Alliance Chemicals. <https://www.alliancechemicals.com/wp-content/uploads/2019/01/MEA-sds.pdf> (accessed 02 February).

[86] Merck. <https://www.sigmaaldrich.com/IN/en/sds/aldrich/471496> (accessed 02 February).

[87] Hemmati A, Rashidi H, Hemmati A, Kazemi A. Using rate based simulation, sensitivity analysis and response surface methodology for optimization of an industrial CO₂

capture plant. J Nat Gas Sci Eng 2019;62:101-12.
<https://doi.org/10.1016/j.jngse.2018.12.002>.

[88] Garcia S, Gil MV, Martín CF, Pis JJ, Rubiera F, Pevida C. Breakthrough adsorption study of a commercial activated carbon for pre-combustion CO₂ capture. Chem Eng J 2011;171:549-56. <https://doi.org/10.1016/j.cej.2011.04.027>.

[89] Shafeeyan MS, Daud WM, Houshmand A, Arami-Niya A. The application of response surface methodology to optimize the amination of activated carbon for the preparation of carbon dioxide adsorbents. Fuel 2012;94:465-72.
<https://doi.org/10.1016/j.fuel.2011.11.035>.

[90] Gupta AK, Gautam A, Mondal MK. Experimental, modeling and RSM optimization of CO₂ loading for an aqueous blend of diethylenetriamine and 3-dimethyl amino-1-propanol. Korean J Chem Eng 2023;40:1151-67. <https://doi.org/10.1007/s11814-022-1300-3>

References

- [1] Gautam A, Mondal MK. Post-combustion capture of CO₂ using novel aqueous triethylenetetramine and 2-dimethylaminoethanol amine blend: equilibrium CO₂ loading-empirical model and optimization, CO₂ desorption, absorption heat, and ¹³C NMR analysis. *Fuel*. 2023;331:125864. <https://doi.org/10.1016/j.fuel.2022.125864>
- [2] Gautam A, Mondal MK. Novel aqueous amine blend of 2-(Butylamino) ethanol and 2-Dimethylaminoethanol for CO₂ capture: Equilibrium CO₂ loading, RSM optimization, desorption study, characterization and toxicity assessment. *Sep Purif Technol*. 2023;322:124279. <https://doi.org/10.1016/j.seppur.2023.124279>
- [3] Gautam A, Mondal MK. Review of recent trends and various techniques for CO₂ capture: Special emphasis on biphasic amine solvents. *Fuel*. 2023;334:126616. <https://doi.org/10.1016/j.fuel.2022.126616>
- [4] Zhang R, Li Y, He X, Niu Y, Li CE, Amer MW, Barzagli F. Investigation of the improvement of the CO₂ capture performance of aqueous amine sorbents by switching from dual-amine to trio-amine systems. *Sep Purif Technol*. 2023;316:123810. <https://doi.org/10.1016/j.seppur.2023.123810>
- [5] Agnihotri N, Gupta GK, Mondal MK. Thermo-kinetic analysis, thermodynamic parameters and comprehensive pyrolysis index of Melia azedarach sawdust as a genesis of bioenergy. *Biomass Conv Bioref* 2022;1-8. <https://doi.org/10.1007/s13399-022-02524-y>
- [6] He X, He H, Barzagli F, Amer MW, Li CE, Zhang R. Analysis of the energy consumption in solvent regeneration processes using binary amine blends for CO₂ capture. *Energy*. 2023;270:126903. <https://doi.org/10.1016/j.energy.2023.126903>

- [7] Zhao S, Wang Y, Zhu K, Zhao D, Song Q. Improving the absorption load, high viscosity, and regeneration efficiency of CO₂ capture using a novel tri-solvent biphasic solvents of TETA-AMP-1DMA2P. *Environ Sci Pollut Res* 2022;29:84903-15. <https://doi.org/10.1007/s11356-022-21822-6>
- [8] Pandey D, Mondal MK. Thermodynamic modeling and new experimental CO₂ solubility into aqueous EAE and AEEA blend, heat of absorption, cyclic absorption capacity and desorption study for post-combustion CO₂ capture. *Chem Eng J*. 2021;410:128334. <https://doi.org/10.1016/j.cej.2020.128334>
- [9] Pandey D, Mondal MK. Equilibrium CO₂ solubility in the aqueous mixture of MAE and AEEA: Experimental study and development of modified thermodynamic model. *Fluid Ph Equilibria*. 2020;522:112766. <https://doi.org/10.1016/j.fluid.2020.112766>
- [10] Gupta AK, Gautam A, Mondal MK. Experimental, modeling and RSM optimization of CO₂ loading for an aqueous blend of diethylenetriamine and 3-dimethyl amino-1-propanol. *Korean J Chem Eng* 2023;40:1151-67. <https://doi.org/10.1007/s11814-022-1300-3>
- [11] Lv B, Guo B, Zhou Z, Jing G. Mechanisms of CO₂ capture into monoethanolamine solution with different CO₂ loading during the absorption/desorption processes. *Environ Sci Technol* 2015;49:10728-35. <https://doi.org/10.1021/acs.est.5b02356>
- [12] Mazari SA, Kang TH, Devkota S, Cha JY, Shin BJ, Mun JH, Kim KM, Lee U, Moon JH. Investigating the effect of blending of diamine and alkanolamine for CO₂ capture: Experiment and thermodynamic modeling of CO₂-AEEA-DEA-H₂O system. *Chem Eng J*. 2023;470:144141. <https://doi.org/10.1016/j.cej.2023.144141>
- [13] Kumari M, Vega F, Fernández LM, Shadangi KP, Kumar N. Liquid Amine functional, aqueous blends and the CO₂ absorption capacity: molecular structure, size, interaction

- parameter and mechanistic aspects. *J Mol Liq.* 2023;384:122288. <https://doi.org/10.1016/j.molliq.2023.122288>
- [14] Perumal M, Jayaraman D, Balraj A. Experimental studies on CO₂ absorption and solvent recovery in aqueous blends of monoethanolamine and tetrabutylammonium hydroxide. *Chemosphere.* 2021;276:130159. <https://doi.org/10.1016/j.chemosphere.2021.130159>
- [15] Kumar S, Mondal MK. Selection of efficient absorbent for CO₂ capture from gases containing low CO₂. *Korean J Chem Eng* 2020;37:231-9. <https://doi.org/10.1007/s11814-019-0440-6>
- [16] Kim J, Lee J, Lee Y, Kim H, Kim E, Lee KS. Evaluation of aqueous polyamines as CO₂ capture solvents. *Energy.* 2019;187:115908. <https://doi.org/10.1016/j.energy.2019.115908>
- [17] Ji L, Yu H, Li K, Yu B, Grigore M, Yang Q, Wang X, Chen Z, Zeng M, Zhao S. Integrated absorption-mineralisation for low-energy CO₂ capture and sequestration. *Appl Energy* 2018;225: 356-66. <https://doi.org/10.1016/j.apenergy.2018.04.108>
- [18] Bains P, Psarras P, Wilcox J. CO₂ capture from the industry sector. *Prog Energy Combust Sci* 2017;63:146-72. <https://doi.org/10.1016/j.pecs.2017.07.001>
- [19] Gao G, Jiang W, Li X, Zhao Z, Jiang C, Luo C, Wu F, Zhang L. Novel assessment of highly efficient polyamines for post-combustion CO₂ capture: Absorption heat, reaction rate, CO₂ cyclic capacity, and phase change behavior. *Sep Purif Technol.* 2023;306:122615. <https://doi.org/10.1016/j.seppur.2022.122615>
- [20] Li T, Yang C, Tantikhajorngosol P, Sema T, Shi H, Tontiwachwuthikul P. Experimental investigations and the modeling approach for CO₂ solubility in aqueous

blended amine systems of monoethanolamine, 2-amino-2-methyl-1-propanol, and 2-(butylamino) ethanol. *Environ Sci Pollut Res* 2022;29:69402-23. <https://doi.org/10.1007/s11356-022-20411-x>

[21] Shen Y, Jiang C, Zhang S, Chen J, Wang L, Chen J. Biphasic solvent for CO₂ capture: Amine property-performance and heat duty relationship. *Appl Energy* 2018;230:726-33. <https://doi.org/10.1016/j.apenergy.2018.09.005>

[22] Zhang S, Shen Y, Shao P, Chen J, Wang L. Kinetics, thermodynamics, and mechanism of a novel biphasic solvent for CO₂ capture from flue gas. *Environ Sci Technol* 2018;52:3660-8. <https://doi.org/10.1021/acs.est.7b05936>

[23] Sreedhar I, Nahar T, Venugopal A, Srinivas B. Carbon capture by absorption—Path covered and ahead. *Renew Sustain Energy Rev* 2017;76:1080-107. <https://doi.org/10.1016/j.rser.2017.03.109>

[24] Muchan P, Saiwan C, Narku-Tetteh J, Idem R, Supap T, Tontiwachwuthikul P. Screening tests of aqueous alkanolamine solutions based on primary, secondary, and tertiary structure for blended aqueous amine solution selection in post combustion CO₂ capture. *Chem Eng Sci* 2017;170:574-82. <https://doi.org/10.1016/j.ces.2017.02.031>

[25] Ping T, Dong Y, Shen S. Densities, viscosities and spectroscopic study of partially CO₂-loaded nonaqueous blends of 2-butoxyethanol with 2-(ethylamino) ethanol and 2-(butylamino) ethanol at temperatures of (293.15 to 353.15) K. *J Mol Liq*. 2020;312:113389. <https://doi.org/10.1016/j.molliq.2020.113389>

[26] Arshad N, Alhajaj A. Process synthesis for amine-based CO₂ capture from combined cycle gas turbine power plant. *Energy*. 2023;274:127391. <https://doi.org/10.1016/j.energy.2023.127391>

- [27] Singh S, Pandey D, Mondal MK. New experimental data on equilibrium CO₂ loading into aqueous 3-dimethyl amino-1-propanol and 1, 5-diamino-2-methylpentane blend: Empirical model and CO₂ absorption enthalpy. *J Chem Eng Data* 2020;66:740-8. <https://doi.org/10.1021/acs.jced.0c00851>
- [28] Hafizi A, Mokari MH, Khalifeh R, Farsi M, Rahimpour MR. Improving the CO₂ solubility in aqueous mixture of MDEA and different polyamine promoters: The effects of primary and secondary functional groups. *J Mol Liq.* 2020;297:111803. <https://doi.org/10.1016/j.molliq.2019.111803>
- [29] Kim YE, Yun SH, Choi JH, Nam SC, Park SY, Jeong SK, Yoon YI. Comparison of the CO₂ absorption characteristics of aqueous solutions of diamines: absorption capacity, specific heat capacity, and heat of absorption. *Energy Fuels* 2015;29:2582-90. <https://doi.org/10.1021/ef500561a>
- [30] Yang C, Li T, Tantikhajongsol P, Sema T, Xiao M, Tontiwachwuthikul P. Evaluation of novel aqueous piperazine-based physical-chemical solutions as biphasic solvents for CO₂ capture: Initial absorption rate, equilibrium solubility, phase separation and desorption rate. *Chem Eng Sci.* 2023;277:118852. <https://doi.org/10.1016/j.ces.2023.118852>
- [31] Wai SK, Nwaoha C, Saiwan C, Idem R, Supap T. Absorption heat, solubility, absorption and desorption rates, cyclic capacity, heat duty, and absorption kinetic modeling of AMP–DETA blend for post–combustion CO₂ capture. *Sep Purif Technol* 2018;194:89-95. <https://doi.org/10.1016/j.seppur.2017.11.024>

- [32] Arshad MW, Von Solms N, Thomsen K. Thermodynamic modeling of liquid–liquid phase change solvents for CO₂ capture. *Int J Greenh Gas Control* 2016;53:401-24. <https://doi.org/10.1016/j.ijggc.2016.08.014>
- [33] Shi H, Naami A, Idem R, Tontiwachwuthikul P. Catalytic and non catalytic solvent regeneration during absorption-based CO₂ capture with single and blended reactive amine solvents. *Int J Greenh Gas Control* 2014;26:39-50. <https://doi.org/10.1016/j.ijggc.2014.04.007>
- [34] Mun JH, Shin BJ, Kim SM, You JK, Park YC, Chun DH, Lee JS, Min BM, Lee U, Kim KM, Moon JH. Optimal MEA/DIPA/water blending ratio for minimizing regeneration energy in absorption-based carbon capture process: Experimental CO₂ solubility and thermodynamic modeling. *Chem Eng J.* 2022;444:136523. <https://doi.org/10.1016/j.cej.2022.136523>
- [35] Choi BK, Kim SM, Lee JS, Park YC, Chun DH, Shin HY, Sung HJ, Min BM, Moon JH. Effect of blending ratio and temperature on CO₂ solubility in blended aqueous solution of monoethanolamine and 2-amino-2-methyl-propanol: experimental and modeling study using the electrolyte nonrandom two-liquid model. *ACS omega* 2020;5:28738-48. <https://doi.org/10.1021/acsomega.0c04046>
- [36] Ping T, Dong Y, Shen S. Energy-efficient CO₂ capture using nonaqueous absorbents of secondary alkanolamines with a 2-butoxyethanol cosolvent. *ACS Sustain Chem Eng* 2020;8: 18071-82. <https://doi.org/10.1021/acssuschemeng.0c06345>
- [37] Mondal BK, Bandyopadhyay SS, Samanta AN. Kinetics of CO₂ absorption in aqueous hexamethylenediamine. *Int J Greenh Gas Control* 2017;56:116-25. <https://doi.org/10.1016/j.ijggc.2016.11.023>

- [38] Vaidya PD, Kenig EY. CO₂-alkanolamine reaction kinetics: a review of recent studies. *Chem Eng Technol* 2007;30:1467-74. <https://doi.org/10.1002/ceat.200700268>
- [39] Hu X, Huang J, He X, Luo Q, Li CE, Zhou C, Zhang R. Analyzing the potential benefits of trio-amine systems for enhancing the CO₂ desorption processes. *Fuel*. 2022;316:123216. <https://doi.org/10.1016/j.fuel.2022.123216>
- [40] Chen M, Li M, Zhang F, Hu X, Wu Y. Fast and efficient CO₂ absorption in non-aqueous tertiary amines promoted by ethylene glycol. *Energy Fuels* 2022;36:4830-6. <https://doi.org/10.1021/acs.energyfuels.2c00215>
- [41] Kim YE, Moon SJ, Yoon YI, Jeong SK, Park KT, Bae ST, Nam SC. Heat of absorption and absorption capacity of CO₂ in aqueous solutions of amine containing multiple amino groups. *Sep Purif Technol* 2014;122:112-8. <https://doi.org/10.1016/j.seppur.2013.10.030>
- [42] Sanz-Pérez ES, Arencibia A, Sanz R, Calleja G. New developments on carbon dioxide capture using amine-impregnated silicas. *Adsorption* 2016;22:609-19. <https://doi.org/10.1007/s10450-015-9740-2>
- [43] Xiao M, Liu H, Gao H, Liang Z. CO₂ absorption with aqueous tertiary amine solutions: Equilibrium solubility and thermodynamic modeling. *J Chem Thermodyn* 2018;122:170-82. <https://doi.org/10.1016/j.jct.2018.03.020>
- [44] Narku-Tetteh J, Muchan P, Saiwan C, Supap T, Idem R. Selection of components for formulation of amine blends for post combustion CO₂ capture based on the side chain structure of primary, secondary and tertiary amines. *Chem Eng Sci* 2017;170:542-60. <https://doi.org/10.1016/j.ces.2017.02.036>

- [45] Cao F, Gao H, Xiong Q, Liang Z. Experimental studies on mass transfer performance for CO₂ absorption into aqueous N, N-dimethylethanolamine (DMEA) based solutions in a PTFE hollow fiber membrane contactor. *Int J Greenh Gas Control* 2019;82:210-7. <https://doi.org/10.1016/j.ijggc.2018.12.011>
- [46] El Hadri N, Quang DV, Goetheer EL, Zahra MR. Aqueous amine solution characterization for post-combustion CO₂ capture process. *Appl Energy* 2017;185:1433-49. <https://doi.org/10.1016/j.apenergy.2016.03.043>
- [47] Mondal BK, Bandyopadhyay SS, Samanta AN. Vapor–liquid equilibrium measurement and ENRTL modeling of CO₂ absorption in aqueous hexamethylenediamine. *Fluid Ph Equilibria* 2015;402:102-12. <https://doi.org/10.1016/j.fluid.2015.05.033>
- [48] Singh P, van Swaaij WP, Brillman DW. Kinetics study of carbon dioxide absorption in aqueous solutions of 1, 6-hexamethyldiamine (HMDA) and 1, 6-hexamethyldiamine, N, N' di-methyl (HMDA, N, N'). *Chem Eng Sci* 2011;66:4521-32. <https://doi.org/10.1016/j.ces.2011.06.008>
- [49] Pasha M, Li G, Shang M, Liu S, Su Y. Mass transfer and kinetic characteristics for CO₂ absorption in microstructured reactors using an aqueous mixed amine. *Sep Purif Technol.* 2021;274:118987. <https://doi.org/10.1016/j.seppur.2021.118987>
- [50] Patil MP, Vaidya PD. Kinetics of CO₂ absorption into aqueous AMP/HMDA/TEG mixtures. *ChemistrySelect* 2018;3:195-200. <https://doi.org/10.1002/slct.201702464>
- [51] Azhgan M, Farsi M, Eslamloueyan R. Solubility of CO₂ in aqueous solutions of DAMP+MDEA, DAMP+MEA, DAH+MDEA and DAH+MEA. *J Nat Gas Sci Eng* 2017;46: 526-32. <https://doi.org/10.1016/j.jngse.2017.08.014>

- [52] El Hadri N, Quang DV, Goetheer EL, Zahra MR. Aqueous amine solution characterization for post-combustion CO₂ capture process. *Appl Energy* 2017;185:1433-49. <https://doi.org/10.1016/j.apenergy.2016.03.043>
- [53] Ling H, Gao H, Liang Z. Comprehensive solubility of N₂O and mass transfer studies on an effective reactive N, N-dimethylethanolamine (DMEA) solvent for post-combustion CO₂ capture. *Chem Eng J* 2019;355:369-79. <https://doi.org/10.1016/j.cej.2018.08.147>
- [54] Mokhtari Z, Pakravesch A, Zarei H. High-pressure densities of 2-(dimethylamino) ethanol and 2-(diethylamino) ethanol: Measurement and modeling with new modified Tait and PC-SAFT equations of state. *Fluid Ph Equilibria*. 2023;572:113825. <https://doi.org/10.1016/j.fluid.2023.113825>
- [55] Sharif M, Zhang T, Wu X, Yu Y, Zhang Z. Evaluation of CO₂ absorption performance by molecular dynamic simulation for mixed secondary and tertiary amines. *Int J Greenh Gas Control*. 2020;97:103059. <https://doi.org/10.1016/j.ijggc.2020.103059>
- [56] Sharif M, Fan H, Wu X, Yu Y, Zhang T, Zhang Z. Assessment of novel solvent system for CO₂ capture applications. *Fuel*. 2023;337:127218. <https://doi.org/10.1016/j.fuel.2022.127218>
- [57] Zhang P, Xu R, Li H, Gao H, Liang Z. Mass transfer performance for CO₂ absorption into aqueous blended DMEA/MEA solution with optimized molar ratio in a hollow fiber membrane contactor. *Sep Purif Technol* 2019;211:628-36. <https://doi.org/10.1016/j.seppur.2018.10.034>
- [58] Ling H, Liu S, Wang T, Gao H, Liang Z. Characterization and correlations of CO₂ absorption performance into aqueous amine blended solution of monoethanolamine (MEA)

- and N, N-dimethylethanolamine (DMEA) in a packed column. *Energy Fuels* 2019;33:7614-25. <https://doi.org/10.1021/acs.energyfuels.9b01764>
- [59] Jiang W, Luo X, Gao H, Liang Z, Liu B, Tontiwachwuthikul P, Hu X. A comparative kinetics study of CO₂ absorption into aqueous DEEA/MEA and DMEA/MEA blended solutions. *AIChE J* 2018;64:1350-8. <https://doi.org/10.1002/aic.16024>
- [60] Chowdhury FA, Okabe H, Shimizu S, Onoda M, Fujioka Y. Development of novel tertiary amine absorbents for CO₂ capture. *Energy Procedia* 2009;1:1241-8. <https://doi.org/10.1016/j.egypro.2009.01.163>
- [61] Buvik V, Vevelstad SJ, Brakstad OG, Knuutila HK. Stability of structurally varied aqueous amines for CO₂ capture. *Ind Eng Chem Res* 2021;60:5627-38. <https://doi.org/10.1021/acs.iecr.1c00502>
- [62] Brúder P, Lauritsen KG, Mejdell T, Svendsen HF. CO₂ capture into aqueous solutions of 3-methylaminopropylamine activated dimethyl-monoethanolamine. *Chem Eng Sci* 2012;75:28-37. <https://doi.org/10.1016/j.ces.2012.03.005>
- [63] Agnihotri N, Mondal MK. Process parameter variation of Melia azedarach sawdust pyrolysis for fuel properties, physicochemical characterization, and in-depth speciation analysis. *Biomass Conv Bioref* 2023:1-5. <https://doi.org/10.1007/s13399-023-04305-7>
- [64] Zheng W, Yan Z, Zhang R, Jiang W, Luo X, Liang Z, Yang Q, Yu H. A study of kinetics, equilibrium solubility, speciation and thermodynamics of CO₂ absorption into benzylamine (BZA) solution. *Chem Eng Sci*. 2022;251:117452. <https://doi.org/10.1016/j.ces.2022.117452>

- [65] Huang Q, Jing G, Zhou X, Lv B, Zhou Z. A novel biphasic solvent of amino-functionalized ionic liquid for CO₂ capture: High efficiency and regenerability. *J CO₂ Util* 2018;25:22-30. <https://doi.org/10.1016/j.jcou.2018.03.001>
- [66] Zhou X, Liu F, Lv B, Zhou Z, Jing G. Evaluation of the novel biphasic solvents for CO₂ capture: Performance and mechanism. *Int J Greenh Gas Control* 2017;60:120-8. <https://doi.org/10.1016/j.ijggc.2017.03.013>
- [67] Perinu C, Arstad B, Jens KJ. NMR spectroscopy applied to amine-CO₂-H₂O systems relevant for post-combustion CO₂ capture: A review. *Int J Greenh Gas Control* 2014;20:230-43. <https://doi.org/10.1016/j.ijggc.2013.10.029>
- [68] Shafaghat J, Ghaemi A. Comparison of Pb (II) adsorption by ground granulated blast-furnace and phosphorus slags; exploitation of RSM. *Iran J Sci Technol Trans Sci* 2021;45:899-911. <https://doi.org/10.1007/s40995-021-01075-7>
- [69] Asgarifard P, Rahimi M, Tafreshi N. Response surface modelling of CO₂ capture by ammonia aqueous solution in a microchannel. *Can J Chem Eng* 2021;99:601-12. <https://doi.org/10.1002/cjce.23881>
- [70] Pashaei H, Ghaemi A, Nasiri M, Karami B. Experimental modeling and optimization of CO₂ absorption into piperazine solutions using RSM-CCD methodology. *ACS Omega* 2020;5:8432-48. <https://doi.org/10.1021/acsomega.9b03363>
- [71] Shafeeyan MS, Daud WM, Houshmand A, Arami-Niya A. The application of response surface methodology to optimize the amination of activated carbon for the preparation of carbon dioxide adsorbents. *Fuel* 2012;94:465-72. <https://doi.org/10.1016/j.fuel.2011.11.035>
- [72] Sibuh BZ, Gupta PK, Taneja P, Khanna S, Sarkar P, Pachisia S, Khan AA, Jha NK, Dua K, Singh SK, Pandey S. Synthesis, in silico study, and anti-cancer activity of

thiosemicarbazone derivatives. Biomedicines. 2021;9:1375.

<https://doi.org/10.3390/biomedicines9101375>

[73] Gadaleta D, Vuković K, Toma C, Lavado GJ, Karmaus AL, Mansouri K, Kleinstreuer NC, Benfenati E, Roncaglioni A. SAR and QSAR modeling of a large collection of LD₅₀ rat acute oral toxicity data. J Cheminformatics 2019;11:1-6.

<https://doi.org/10.1186/s13321-019-0383-2>

[74] Strickland J, Clippinger AJ, Brown J, Allen D, Jacobs A, Matheson J, Lowit A, Reinke EN, Johnson MS, Quinn Jr MJ, Mattie D. Status of acute systemic toxicity testing requirements and data uses by US regulatory agencies. Regul Toxicol Pharmacol 2018;94:183-96. <https://doi.org/10.1016/j.yrtph.2018.01.022>

[75] Li X, Chen L, Cheng F, Wu Z, Bian H, Xu C, Li W, Liu G, Shen X, Tang Y. In silico prediction of chemical acute oral toxicity using multi-classification methods. J Chem Inf Model 2014;54:1061-9. <https://doi.org/10.1021/ci5000467>

[76] Wai SK, Saiwan C, Idem R, Supap T, Nwaoha C. Carbon Dioxide (CO₂) Solubility in Diethylenetriamine and 2-Amino-2-Methyl-1-Propanal (DETA–AMP) Solvent System for Amine–Based CO₂ Capture in Flue Gas from Coal Combustion. Energy Procedia 2017;114: 1973-9. <https://doi.org/10.1016/j.egypro.2017.03.1329>

[77] Ramezani R, Mazinani S, Di Felice R. Density, Viscosity, pH, Heat of Absorption, and CO₂ Loading Capacity of Methyldiethanolamine and Potassium Lysinate Blend Solutions. J Chem Eng Data 2021;66:1611-29. <https://doi.org/10.1021/acs.jced.0c00855>

[78] Ahmed RE, Wiheeb AD. Enhancement of carbon dioxide absorption into aqueous potassium carbonate by adding amino acid salts. Mater Today: Proc 2020;20:611-6. <https://doi.org/10.1016/j.matpr.2019.09.198>

- [79] Pandey D, Mondal MK. Experimental data and modeling for density and viscosity of carbon dioxide (CO₂)-loaded and-unloaded aqueous blend of 2-(ethylamino) ethanol (EAE) and aminoethylethanolamine (AEEA) for post-combustion CO₂ capture. *J Mol Liq.* 2021;330: 115678. <https://doi.org/10.1016/j.molliq.2021.115678>
- [80] Shen S, Yang YN, Wang Y, Ren S, Han J, Chen A. CO₂ absorption into aqueous potassium salts of lysine and proline: density, viscosity and solubility of CO₂. *Fluid Ph Equilibria* 2015;399: 40-9. <https://doi.org/10.1016/j.fluid.2015.04.021>
- [81] Sobrino M, Concepción EI, Gómez-Hernández Á, Martín MC, Segovia JJ. Viscosity and density measurements of aqueous amines at high pressures: MDEA-water and MEA-water mixtures for CO₂ capture. *J Chem Thermodyn* 2016;98:231-41. <https://doi.org/10.1016/j.jct.2016.03.021>
- [82] Hsu CH, Li MH. Densities of aqueous blended amines. *J Chem Eng Data* 1997;42:502-7. <https://doi.org/10.1021/je960356j>
- [83] Böttinger W, Maiwald M, Hasse H. Online NMR Spectroscopic Study of Species Distribution in MDEA–H₂O–CO₂ and MDEA–PIP–H₂O–CO₂. *Ind Eng Chem Res* 2008;47: 7917-26. <https://doi.org/10.1021/ie800914m>
- [84] Kortunov PV, Siskin M, Baugh LS, Calabro DC. In situ nuclear magnetic resonance mechanistic studies of carbon dioxide reactions with liquid amines in aqueous systems: New insights on carbon capture reaction pathways. *Energy Fuels* 2015;29:5919-39. <https://doi.org/10.1021/acs.energyfuels.5b00850>
- [85] Guo C, Chen S, Zhang Y. A ¹³C NMR study of carbon dioxide absorption and desorption in pure and blended 2-(2-aminoethylamine) ethanol (AEEA) and 2-amino-2-

- methyl-1-propanol (AMP) solutions. *Int J Greenh Gas Control* 2014;28:88-95.
<https://doi.org/10.1016/j.ijggc.2014.06.025>
- [86] Agnihotri N, Mondal MK. Comparison of non-catalytic and in-situ catalytic pyrolysis of *Melia azedarach* sawdust. *J Anal Appl Pyrolysis*. 2023;172:106006.
<https://doi.org/10.1016/j.jaap.2023.106006>
- [87] Nandiyanto AB, Oktiani R, Ragadhita R. How to read and interpret FTIR spectroscopy of organic material. *Indones J Sci Technol* 2019;4:97-118.
<https://doi.org/10.17509/ijost.v4i1.15806>
- [88] Gao G, Li X, Jiang W, Zhao Z, Xu Y, Wu F, Luo C, Zhang L. Improved quasi-cycle capacity method based on microcalorimetry strategy for the fast screening of amino acid salt absorbents for CO₂ capture. *Sep Purif Technol*. 2022;289:120767.
<https://doi.org/10.1016/j.seppur.2022.120767>
- [89] Xiao M, Cui D, Liu H, Tontiwachwuthikul P, Liang Z. A new model for correlation and prediction of equilibrium CO₂ solubility in N-methyl-4-piperidinol solvent. *AIChE J* 2017;63: 3395-403. <https://doi.org/10.1002/aic.15709>
- [90] Xie Q, Aroonwilas A, Veawab A. Measurement of heat of CO₂ absorption into 2-amino-2-methyl-1-propanol (AMP)/piperazine (PZ) blends using differential reaction calorimeter. *Energy Procedia* 2013;37:826-33.
<https://doi.org/10.1016/j.egypro.2013.05.175>
- [91] Mathias PM, O'Connell JP. The Gibbs–Helmholtz equation and the thermodynamic consistency of chemical absorption data. *Ind Eng Chem Res* 2012;51:5090-7.
<https://doi.org/10.1021/ie202668k>

- [92] Rochelle G, Chen E, Freeman S, Van Wagener D, Xu Q, Voice A. Aqueous piperazine as the new standard for CO₂ capture technology. *Chem Eng J* 2011;171:725-33. <https://doi.org/10.1016/j.cej.2011.02.011>
- [93] Rho SW, Yoo KP, Lee JS, Nam SC, Son JE, Min BM. Solubility of CO₂ in aqueous methyldiethanolamine solutions. *J Chem Eng Data* 1997;42:1161-4. <https://doi.org/10.1021/je970097d>
- [94] Elmobarak WF, Almomani F, Tawalbeh M, Al-Othman A, Martis R, Rasool K. Current status of CO₂ capture with ionic liquids: Development and progress. *Fuel*. 2023;344:128102. <https://doi.org/10.1016/j.fuel.2023.128102>
- [95] Shao R, Stangeland A. Amines used in CO₂ capture-health and environmental impacts. *Bellona report* 2009;49:1-49.
- [96] Hamm J, Allen D, Ceger P, Flint T, Lowit A, O'Dell L, Tao J, Kleinstreuer N. Performance of the GHS mixtures equation for predicting acute oral toxicity. *Regul Toxicol Pharmacol*. 2021;125:105007. <https://doi.org/10.1016/j.yrtph.2021.105007>
- [97] Rey A, Gouedard C, Ledirac N, Cohen M, Dugay J, Vial J, Pichon V, Bertomeu L, Picq D, Bontemps D, Chopin F. Amine degradation in CO₂ capture. 2. New degradation products of MEA. Pyrazine and alkylpyrazines: analysis, mechanism of formation and toxicity. *Int J Greenh Gas Control* 2013;19:576-83. <https://doi.org/10.1016/j.ijggc.2013.10.018>
- [98] Ramezani R, Mazinani S, Di Felice R. State-of-the-art of CO₂ capture with amino acid salt solutions. *Rev Chem Eng* 2022;38:273-99. <https://doi.org/10.1515/revce-2020-0012>
- [99] Caplow M. Kinetics of carbamate formation and breakdown. *J Am Chem Soc* 1968;90: 6795-803. <https://doi.org/10.1021/ja01026a041>

- [100] Danckwerts PV. The reaction of CO₂ with ethanolamines. *Chem Eng Sci* 1979;34:443-6. [https://doi.org/10.1016/0009-2509\(79\)85087-3](https://doi.org/10.1016/0009-2509(79)85087-3)
- [101] Crooks JE, Donnellan JP. Kinetics and mechanism of the reaction between carbon dioxide and amines in aqueous solution. *J Chem Soc Perkin Trans 2* 1989:331-3. <https://doi.org/10.1039/P29890000331>
- [102] Da Silva EF, Svendsen HF. Ab initio study of the reaction of carbamate formation from CO₂ and alkanolamines. *Ind Eng Chem Res* 2004;43:3413-8. <https://doi.org/10.1021/ie030619k>
- [103] Donaldson TL, Nguyen YN. Carbon dioxide reaction kinetics and transport in aqueous amine membranes. *Ind Eng Chem Fundam* 1980;19:260-6. <https://doi.org/10.1021/i160075a005>
- [104] Shen KP, Li MH. Solubility of carbon dioxide in aqueous mixtures of monoethanolamine with methyldiethanolamine. *J Chem Eng Data* 1992;37:96-100. <https://doi.org/10.1021/je00005a025>
- [105] Aronu UE, Gondal S, Hessen ET, Haug-Warberg T, Hartono A, Hoff KA, Svendsen HF. Solubility of CO₂ in 15, 30, 45 and 60 mass% MEA from 40 to 120 C and model representation using the extended UNIQUAC framework. *Chem Eng Sci* 2011;66:6393-406. <https://doi.org/10.1016/j.ces.2011.08.042>
- [106] Gao H, Wu Z, Liu H, Luo X, Liang Z. Experimental studies on the effect of tertiary amine promoters in aqueous monoethanolamine (MEA) solutions on the absorption/stripping performances in post-combustion CO₂ capture. *Energy Fuels* 2017;31:13883-91. <https://doi.org/10.1021/acs.energyfuels.7b02390>

- [107] Tong D, Trusler JM, Maitland GC, Gibbins J, Fennell PS. Solubility of carbon dioxide in aqueous solution of monoethanolamine or 2-amino-2-methyl-1-propanol: Experimental measurements and modeling. *Int J Greenh Gas Control* 2012;6:37-47. <https://doi.org/10.1016/j.ijggc.2011.11.005>
- [108] Lee JL, Otto FD, Mather AE. Equilibrium between carbon dioxide and aqueous monoethanolamine solutions. *J Appl Chem Biotechnol* 1976;26:541-9. <https://doi.org/10.1002/jctb.5020260177>
- [109] Xiao M, Cui D, Liu H, Tontiwachwuthikul P, Liang Z. A new model for correlation and prediction of equilibrium CO₂ solubility in N-methyl-4-piperidinol solvent. *AIChE J* 2017;63: 3395-403. <https://doi.org/10.1002/aic.15709>
- [110] Bajpai A, Mondal MK. Equilibrium solubility of CO₂ in aqueous mixtures of DEA and AEEA. *J Chem Eng Data* 2013;58:1490-5. <https://doi.org/10.1021/je3011776>
- [111] Kumar S, Mondal MK. Equilibrium solubility of CO₂ in aqueous binary mixture of 2-(diethylamine) ethanol and 1, 6-hexamethyldiamine. *Korean J Chem Eng* 2018;35:1335-40. <https://doi.org/10.1007/s11814-018-0044-6>
- [112] Kumar S, Padhan R, Mondal MK. Equilibrium solubility measurement and modeling of CO₂ absorption in aqueous blend of 2-(diethyl amino) ethanol and ethylenediamine. *J Chem Eng Data* 2020;65:523-31. <https://doi.org/10.1021/acs.jced.9b00699>
- [113] Srisang W, Pouryousefi F, Osei PA, Decardi-Nelson B, Akachuku A, Tontiwachwuthikul P, Idem R. Evaluation of the heat duty of catalyst-aided amine-based post combustion CO₂ capture. *Chem Eng Sci* 2017;170:48-57. <https://doi.org/10.1016/j.ces.2017.01.049>

- [114] Hartono A, Mba EO, Svendsen HF. Physical properties of partially CO₂ loaded aqueous monoethanolamine (MEA). *J Chem Eng Data* 2014;59:1808-16. <https://doi.org/10.1021/je401081e>
- [115] Choubtashani S, Rashidi H. CO₂ capture process intensification of water-lean methyl diethanolamine-piperazine solvent: Experiments and response surface modeling. *Energy*. 2023;267:126447. <https://doi.org/10.1016/j.energy.2022.126447>
- [116] Kortunov PV, Siskin M, Baugh LS, Calabro DC. In situ nuclear magnetic resonance mechanistic studies of carbon dioxide reactions with liquid amines in aqueous systems: New insights on carbon capture reaction pathways. *Energy Fuels* 2015;29:5919-39. <https://doi.org/10.1021/acs.energyfuels.5b00850>
- [117] Wang W, Zheng K, Peng Z, Liu B, Jia X, Tian J. Exploiting proton masking to protect amino achieve efficient capture CO₂ by amino-acids deep eutectic solvents. *Sep Purif Technol*. 2022;299:121787. <https://doi.org/10.1016/j.seppur.2022.121787>
- [118] Xu S, Wang YW, Otto FD, Mather AE. Kinetics of the reaction of carbon dioxide with 2-amino-2-methyl-1-propanol solutions, *Chem Eng Sci* 1996;51:841-50. [https://doi.org/10.1016/0009-2509\(95\)00327-4](https://doi.org/10.1016/0009-2509(95)00327-4)
- [119] Amiri M, Shahhosseini S, Ghaemi A. Optimization of CO₂ capture process from simulated flue gas by dry regenerable alkali metal carbonate based adsorbent using response surface methodology. *Energy Fuels* 2017;31:5286-96. <https://doi.org/10.1021/acs.energyfuels.6b03303>
- [120] Gilmour SG. Response surface designs for experiments in bioprocessing. *Biometrics* 2006;62:323-31. <https://doi.org/10.1111/j.1541-0420.2005.00444.x>

[121] Song C, Kitamura Y, Li S. Optimization of a novel cryogenic CO₂ capture process by response surface methodology (RSM). *J Taiwan Inst Chem Eng* 2014;45:1666-76. <https://doi.org/10.1016/j.jtice.2013.12.009>

[122] Gil MV, Martínez M, Garcia S, Rubiera F, Pis JJ, Pevida C. Response surface methodology as an efficient tool for optimizing carbon adsorbents for CO₂ capture. *Fuel Process Technol* 2013;106:55-61. <https://doi.org/10.1016/j.fuproc.2012.06.018>

[123] Saeidi M, Ghaemi A, Tahvildari K, Derakhshi P. Exploiting response surface methodology (RSM) as a novel approach for the optimization of carbon dioxide adsorption by dry sodium hydroxide. *J Chin Chem Soc* 2018;65:1465-75. <https://doi.org/10.1002/jccs.201800012>

[124] Tang Q, Lau YB, Hu S, Yan W, Yang Y, Chen T. Response surface methodology using Gaussian processes: Towards optimizing the trans-stilbene epoxidation over Co²⁺-NaX catalysts. *Chem Eng J* 2010;156:423-31. <https://doi.org/10.1016/j.cej.2009.11.002>

[125]

<https://www.fishersci.com/store/msds?partNumber=AC126720010&productDescription=N=METHYLDIETHANOLAMINE%2C+1KG&vendorId=VN00032119&countryCode=US&language=en> (accessed 09 November 2023).

[126]

<https://www.fishersci.com/store/msds?partNumber=AC115880050&productDescription=3-DIMETHYLAMINO-1-ROPANOL+5G&vendorId=VN00032119&countryCode=US&language=en> (accessed 09 November 2023).

- [127] Alliance Chemicals. <https://www.alliancechemicals.com/wp-content/uploads/2019/01/MEA-sds.pdf> (accessed 09 November 2023).
- [128] Ted Pella. https://www.tedpella.com/SDS_html/18315_sds.pdf (accessed 09 November 2023).
- [129] Sigma-Aldrich. <https://www.sigmaaldrich.com/IN/en/sds/aldrich/h11696> (accessed 09 November 2023).
- [130] Spectrum. https://www.spectrumchemical.com/media/sds/AM156_AGHS.pdf (accessed 09 November 2023).
- [131] Oxford Lab Fine. [https://www.oxfordlabchem.com/msds/PIPERAZINE%20\(Anhydrous\).pdf](https://www.oxfordlabchem.com/msds/PIPERAZINE%20(Anhydrous).pdf) (accessed 09 November 2023).
- [132] Alliance Chemicals. <https://www.alliancechemicals.com/wp-content/uploads/2011/10/DEA-sds.pdf> (accessed 09 November 2023).
- [133] North Metal and Chemical. <https://northmetal.net/wp-content/uploads/Diethylaminoethanol-Diethylethanolamine-DEAE-DEEA-Pennad-150C6H15NO-100-37-8-SDS.pdf> (accessed 09 November 2023).
- [134] Merck. <https://www.sigmaaldrich.com/IN/en/sds/aldrich/471496> (accessed 09 November 2023).

References

- [1] Muchan P, Saiwan C, Narku-Tetteh J, Idem R, Supap T, Tontiwachwuthikul P. Screening tests of aqueous alkanolamine solutions based on primary, secondary, and tertiary structure for blended aqueous amine solution selection in post combustion CO₂ capture. *Chem Eng Sci* 2017;170:574-82. <https://doi.org/10.1016/j.ces.2017.02.031>
- [2] Pandey D, Mondal MK. Thermodynamic modeling and new experimental CO₂ solubility into aqueous EAE and AEEA blend, heat of absorption, cyclic absorption capacity and desorption study for post-combustion CO₂ capture. *Chem Eng J*. 2021;410:128334. <https://doi.org/10.1016/j.cej.2020.128334>
- [3] Hoyos LA, Paviotti MA, Faroldi BM, Cornaglia LM. Coupling of CO₂ capture and methanation processes using catalysts based on silica recovered from rice husks. *Fuel*. 2022;324:124604. <https://doi.org/10.1016/j.fuel.2022.124604>
- [4] Chowdhury FA, Okabe H, Shimizu S, Onoda M, Fujioka Y. Development of novel tertiary amine absorbents for CO₂ capture. *Energy Procedia* 2009;1:1241-8. <https://doi.org/10.1016/j.egypro.2009.01.163>
- [5] Ling H, Gao H, Liang Z. Comprehensive solubility of N₂O and mass transfer studies on an effective reactive N, N-dimethylethanolamine (DMEA) solvent for post-combustion CO₂ capture. *Chem Eng J* 2019;355:369-79. <https://doi.org/10.1016/j.cej.2018.08.147>
- [6] Sharif M, Zhang T, Wu X, Yu Y, Zhang Z. Evaluation of CO₂ absorption performance by molecular dynamic simulation for mixed secondary and tertiary amines. *Int J Greenh Gas Control*. 2020;97:103059. <https://doi.org/10.1016/j.ijggc.2020.103059>

- [7] Zhang J, Agar DW, Zhang X, Geuzebroek F. CO₂ absorption in biphasic solvents with enhanced low temperature solvent regeneration. *Energy Procedia* 2011;4:67-74. <https://doi.org/10.1016/j.egypro.2011.01.024>
- [8] Kumar S, Mondal MK. Equilibrium solubility of CO₂ in aqueous blend of 2-(diethylamine) ethanol and 2-(2-aminoethylamine) ethanol. *J Chem Eng Data* 2018;63:1163-9. <https://doi.org/10.1021/acs.jced.7b00544>
- [9] Kumar S, Mondal MK. Equilibrium solubility of CO₂ in aqueous binary mixture of 2-(diethylamine) ethanol and 1, 6-hexamethyldiamine. *Korean J Chem Eng* 2018;35:1335-40. <https://doi.org/10.1007/s11814-018-0044-6>
- [10] Pandey D, Mondal MK. Equilibrium CO₂ solubility in the aqueous mixture of MAE and AEEA: Experimental study and development of modified thermodynamic model. *Fluid Ph Equilibria*. 2020;522:112766. <https://doi.org/10.1016/j.fluid.2020.112766>
- [11] Singh S, Pandey D, Mondal MK. New Experimental Data on Equilibrium CO₂ Loading into Aqueous 3-Dimethyl Amino-1-propanol and 1, 5-Diamino-2-methylpentane Blend: Empirical Model and CO₂ Absorption Enthalpy. *J Chem Eng Data* 2020;66:740-8. <https://doi.org/10.1021/acs.jced.0c00851>
- [12] Sreedhar I, Nahar T, Venugopal A, Srinivas B. Carbon capture by absorption—Path covered and ahead. *Renew Sustain energy rev* 2017;76:1080-107. <https://doi.org/10.1016/j.rser.2017.03.109>
- [13] Kumar S, Mondal MK. Selection of efficient absorbent for CO₂ capture from gases containing low CO₂. *Korean J Chem Eng* 2020;37:231-9. <https://doi.org/10.1007/s11814-019-0440-6>

- [14] Liu F, Fang M, Dong W, Wang T, Xia Z, Wang Q, Luo Z. Carbon dioxide absorption in aqueous alkanolamine blends for biphasic solvents screening and evaluation. *Appl Energy* 2019;233:468-77. <https://doi.org/10.1016/j.apenergy.2018.10.007>
- [15] El Hadri N, Quang DV, Goetheer EL, Zahra MR. Aqueous amine solution characterization for post-combustion CO₂ capture process. *Appl Energy* 2017;185:1433-49. <https://doi.org/10.1016/j.apenergy.2016.03.043>
- [16] Guo H, Li C, Shi X, Li H, Shen S. Non-aqueous amine-based absorbents for energy efficient CO₂ capture. *Appl Energy* 2019;239:725-34. <https://doi.org/10.1016/j.apenergy.2019.02.019>
- [17] Pandey D, Mondal MK. Experimental data and modeling for viscosity and refractive index of aqueous mixtures with 2-(methylamino) ethanol (MAE) and aminoethylethanolamine (AEEA). *J Chem Eng Data* 2019;64:3346-55. <https://doi.org/10.1021/acs.jced.9b00171>
- [18] Ahmed RE, Wiheeb AD. Enhancement of carbon dioxide absorption into aqueous potassium carbonate by adding amino acid salts. *Mater Today: Proc* 2020;20:611-6. <https://doi.org/10.1016/j.matpr.2019.09.198>
- [19] Gervasi J, Dubois L, Thomas D. Screening tests of new hybrid solvents for the post-combustion CO₂ capture process by chemical absorption. *Energy Procedia* 2014;63:1854-62. <https://doi.org/10.1016/j.egypro.2014.11.193>
- [20] Nwaoha C, Saiwan C, Supap T, Idem R, Tontiwachwuthikul P, Rongwong W, Al-Marri MJ, Benamor A. Carbon dioxide (CO₂) capture performance of aqueous tri-solvent blends containing 2-amino-2-methyl-1-propanol (AMP) and methyldiethanolamine

(MDEA) promoted by diethylenetriamine (DETA). *Int J Greenh Gas Control* 2016;53:292-04. <https://doi.org/10.1016/j.ijggc.2016.08.012>

[21] Cao F, Gao H, Xiong Q, Liang Z. Experimental studies on mass transfer performance for CO₂ absorption into aqueous N, N-dimethylethanolamine (DMEA) based solutions in a PTFE hollow fiber membrane contactor. *Int J Greenh Gas Control* 2019;82:210-7. <https://doi.org/10.1016/j.ijggc.2018.12.011>

[22] Zhang P, Xu R, Li H, Gao H, Liang Z. Mass transfer performance for CO₂ absorption into aqueous blended DMEA/MEA solution with optimized molar ratio in a hollow fiber membrane contactor. *Sep Purif Technol* 2019;211:628-36. <https://doi.org/10.1016/j.seppur.2018.10.034>

[23] Shen Y, Chen H, Wang J, Zhang S, Jiang C, Ye J, Wang L, Chen J. Two-stage interaction performance of CO₂ absorption into biphasic solvents: Mechanism analysis, quantum calculation and energy consumption. *Appl Energy*. 2020;260:114343. <https://doi.org/10.1016/j.apenergy.2019.114343>

[24] Zhang J, Qiao Y, Wang W, Misch R, Hussain K, Agar DW. Development of an energy-efficient CO₂ capture process using thermomorphic biphasic solvents. *Energy Procedia* 2013;37:1254-61. <https://doi.org/10.1016/j.egypro.2013.05.224>

[25] Xiao M, Cui D, Liu H, Tontiwachwuthikul P, Liang Z. A new model for correlation and prediction of equilibrium CO₂ solubility in N- methyl- 4- piperidinol solvent. *AIChE J* 2017;63:3395-403. <https://doi.org/10.1002/aic.15709>

[26] Kim YE, Park JH, Yun SH, Nam SC, Jeong SK, Yoon YI. Carbon dioxide absorption using a phase transitional alkanolamine–alcohol mixture. *J Ind Eng Chem* 2014;20:1486-92. <https://doi.org/10.1016/j.jiec.2013.07.036>

- [27] Mofarahi M, Khojasteh Y, Khaledi H, Farahnak A. Design of CO₂ absorption plant for recovery of CO₂ from flue gases of gas turbine. *Energy* 2008;33:1311-9. <https://doi.org/10.1016/j.energy.2008.02.013>
- [28] Conway W, Bruggink S, Beyad Y, Luo W, Melián-Cabrera I, Puxty G, Feron P. CO₂ absorption into aqueous amine blended solutions containing monoethanolamine (MEA), N, N-dimethylethanolamine (DMEA), N, N-diethylethanolamine (DEEA) and 2-amino-2-methyl-1-propanol (AMP) for post-combustion capture processes. *Chem Eng Sci* 2015;126:446-54. <https://doi.org/10.1016/j.ces.2014.12.053>
- [29] Luo X, Fu K, Yang Z, Gao H, Rongwong W, Liang Z, Tontiwachwuthikul P. Experimental studies of reboiler heat duty for CO₂ desorption from triethylenetetramine (TETA) and triethylenetetramine (TETA) + N-methyldiethanolamine (MDEA). *Ind Eng Chem Res* 2015;54:8554-60. <https://doi.org/10.1021/acs.iecr.5b00158>
- [30] Quan C, Chu H, Zhou Y, Su S, Su R, Gao N. Amine-modified silica zeolite from coal gangue for CO₂ capture. *Fuel*. 2022;322:124184. <https://doi.org/10.1016/j.fuel.2022.124184>
- [31] Liu F, Jing G, Zhou X, Lv B, Zhou Z. Performance and mechanisms of triethylene tetramine (TETA) and 2-amino-2-methyl-1-propanol (AMP) in aqueous and non-aqueous solutions for CO₂ capture. *ACS Sustain Chem Eng* 2018;6:1352-61. <https://doi.org/10.1021/acssuschemeng.7b03717>
- [32] Ramezani R, Mazinani S, Di Felice R. Experimental Study of CO Absorption in Potassium Carbonate Solution Promoted by Triethylenetetramine. *Open Chem Eng J* 2018;12:67-79. [10.2174/1874123101812010067](https://doi.org/10.2174/1874123101812010067)

- [33] Schäffer A, Brechtel K, Scheffknecht G. Comparative study on differently concentrated aqueous solutions of MEA and TETA for CO₂ capture from flue gases. *Fuel* 2012;101:148-53. <https://doi.org/10.1016/j.fuel.2011.06.037>
- [34] Tzirakis F, Papadopoulos AI, Seferlis P, Tsivintzelis I. CO₂ Solubility in diethylenetriamine (DETA) and triethylenetetramine (TETA) aqueous mixtures: Experimental investigation and correlation using the CPA equation of state. *Chem Thermody Therm Analysis*. 2021;3:100017. <https://doi.org/10.1016/j.ctta.2021.100017>
- [35] Wang M, Rao N, Liu Y, Li J, Cheng Q, Li J. Enhancement of CO₂ capture in the MDEA solution by introducing TETA or TETA-AEP mixtures as an activator. *Sep Sci Technol* 2019;54:101-9. <https://doi.org/10.1080/01496395.2018.1504797>
- [36] Zheng S, Tao M, Liu Q, Ning L, He Y, Shi Y. Capturing CO₂ into the precipitate of a phase-changing solvent after absorption. *Environ Sci Technol* 2014;48:8905-10. <https://doi.org/10.1021/es501554h>
- [37] Kortunov PV, Siskin M, Baugh LS, Calabro DC. In situ nuclear magnetic resonance mechanistic studies of carbon dioxide reactions with liquid amines in aqueous systems: New insights on carbon capture reaction pathways. *Energy Fuels* 2015;29:5919-39. <https://doi.org/10.1021/acs.energyfuels.5b00850>
- [38] Barzagli F, Mani F, Peruzzini M. Novel water-free biphasic absorbents for efficient CO₂ capture. *Int J Greenh Gas Control* 2017;60:100-9. <https://doi.org/10.1016/j.ijggc.2017.03.010>
- [39] Wang W, Peng Z, Gao H, Sema T, Shi J, Liang Z. New insight and evaluation of secondary Amine/N-butanol biphasic solutions for CO₂ Capture: Equilibrium Solubility,

phase separation behavior, absorption Rate, desorption Rate, energy consumption and ion species. Chem Eng J. 2022;431:133912. <https://doi.org/10.1016/j.cej.2021.133912>

[40] Longeras O, Gautier A, Ballerat-Busserolles K, Andanson JM. Tuning critical solution temperature for CO₂ capture by aqueous solution of amine. J Mol Liq. 2021;343:117628. <https://doi.org/10.1016/j.molliq.2021.117628>

[41] Lv J, Liu S, Ling H, Gao H, Olson W, Li Q, Bairq ZA, Liang Z. Development of a Promising Biphasic Absorbent for Post combustion CO₂ Capture: Sulfolane + 2-(Methylamino) ethanol + H₂O. Ind Eng Chem Res. 2020;59:14496-06. <https://doi.org/10.1021/acs.iecr.0c02389>

[42] Wang R, Liu S, Wang L, Li Q, Zhang S, Chen B, Jiang L, Zhang Y. Superior energy-saving splitter in monoethanolamine-based biphasic solvents for CO₂ capture from coal-fired flue gas. Appl Energy 2019;242:302-10. <https://doi.org/10.1016/j.apenergy.2019.03.138>

[43] Alkhatib III , Galindo A, Vega LF. Systematic study of the effect of the co-solvent on the performance of amine-based solvents for CO₂ capture. Sep Purif Technol. 2022;282:120093. <https://doi.org/10.1016/j.seppur.2021.120093>

[44] Xu Z, Wang S, Chen C. CO₂ absorption by biphasic solvents: Mixtures of 1, 4-Butanediamine and 2-(Diethylamino)-ethanol. Int J Greenh Gas Control 2013;16:107-15. <https://doi.org/10.1016/j.ijggc.2013.03.013>

[45] Liang ZH, Rongwong W, Liu H, Fu K, Gao H, Cao F, Zhang R, Sema T, Henni A, Sumon K, Nath D. Recent progress and new developments in post-combustion carbon-capture technology with amine based solvents. Int J Greenh Gas Control 2015;40:26-54. <https://doi.org/10.1016/j.ijggc.2015.06.017>

- [46] Nuchitprasittichai A, Cremaschi S. Optimization of CO₂ Capture Process with Aqueous Amines- A Comparison of Two Simulation–Optimization Approaches. *Ind Eng Chem Res* 2013;52:10236-43. <https://doi.org/10.1021/ie3029366>
- [47] Garcia S, Gil MV, Martín CF, Pis JJ, Rubiera F, Pevida C. Breakthrough adsorption study of a commercial activated carbon for pre-combustion CO₂ capture. *Chem Eng J* 2011;171:549-56. <https://doi.org/10.1016/j.cej.2011.04.027>
- [48] Tang Q, Lau YB, Hu S, Yan W, Yang Y, Chen T. Response surface methodology using Gaussian processes: Towards optimizing the trans-stilbene epoxidation over Co²⁺-NaX catalysts. *Chem Eng J* 2010;156:423-31. <https://doi.org/10.1016/j.cej.2009.11.002>
- [49] Sahraie S, Rashidi H, Valeh-e-Sheyda P. An optimization framework to investigate the CO₂ capture performance by MEA: Experimental and statistical studies using Box-Behnken design. *Process Saf Environ Prot* 2019;122:161-8. <https://doi.org/10.1016/j.psep.2018.11.026>
- [50] Hemmati A, Rashidi H, Hemmati A, Kazemi A. Using rate based simulation, sensitivity analysis and response surface methodology for optimization of an industrial CO₂ capture plant. *J Nat Gas Sci Eng* 2019;62:101-12. <https://doi.org/10.1016/j.jngse.2018.12.002>
- [51] Nuchitprasittichai A, Cremaschi S. Optimization of CO₂ capture process with aqueous amines using response surface methodology. *Comput Chem Eng* 2011;35:1521-31. <https://doi.org/10.1016/j.compchemeng.2011.03.016>
- [52] Gil MV, Martínez M, Garcia S, Rubiera F, Pis JJ, Pevida C. Response surface methodology as an efficient tool for optimizing carbon adsorbents for CO₂ capture. *Fuel Process Technol* 2013;106:55-61. <https://doi.org/10.1016/j.fuproc.2012.06.018>

- [53] Song C, Kitamura Y, Li S. Optimization of a novel cryogenic CO₂ capture process by response surface methodology (RSM). *J Taiwan Inst Chem Eng* 2014;45:1666-76. <https://doi.org/10.1016/j.jtice.2013.12.009>
- [54] Das D, Meikap BC. Optimization of process condition for the preparation of amine-impregnated activated carbon developed for CO₂ capture and applied to methylene blue adsorption by response surface methodology. *J Environ Sci Health* 2017;52:1164-72. <https://doi.org/10.1080/10934529.2017.1356204>
- [55] Saeidi M, Ghaemi A, Tahvildari K, Derakhshi P. Exploiting response surface methodology (RSM) as a novel approach for the optimization of carbon dioxide adsorption by dry sodium hydroxide. *J Chin Chem Soc* 2018;65:1465-75. <https://doi.org/10.1002/jccs.201800012>
- [56] Asgarifard P, Rahimi M, Tafreshi N. Response surface modelling of CO₂ capture by ammonia aqueous solution in a microchannel. *Can J Chem Eng* 2021;99:601-12. <https://doi.org/10.1002/cjce.23881>
- [57] Gao H, Wu Z, Liu H, Luo X, Liang Z. Experimental studies on the effect of tertiary amine promoters in aqueous monoethanolamine (MEA) solutions on the absorption/stripping performances in post-combustion CO₂ capture. *Energy fuels* 2017;31:13883-91. <https://doi.org/10.1021/acs.energyfuels.7b02390>
- [58] Rho SW, Yoo KP, Lee JS, Nam SC, Son JE, Min BM. Solubility of CO₂ in aqueous methyldiethanolamine solutions. *J Chem Eng Data* 1997;42:1161-4. <https://doi.org/10.1021/je970097d>

- [59] Kim I, Svendsen HF. Heat of absorption of carbon dioxide (CO₂) in monoethanolamine (MEA) and 2-(aminoethyl) ethanolamine (AEEA) solutions. *Ind Eng Chem Res* 2007;46:5803-9. <https://doi.org/10.1021/ie0616489>
- [60] Mathias PM, O'Connell JP. The Gibbs–Helmholtz equation and the thermodynamic consistency of chemical absorption data. *Ind Eng Chem Res* 2012;51:5090-7. <https://doi.org/10.1021/ie202668k>
- [61] Caplow M. Kinetics of carbamate formation and breakdown. *J Am Chem Soc* 1968;90:6795-803. <https://doi.org/10.1021/ja01026a041>
- [62] Danckwerts PV. The reaction of CO₂ with ethanolamines. *Chem Eng Sci* 1979;34:443-6. [https://doi.org/10.1016/0009-2509\(79\)85087-3](https://doi.org/10.1016/0009-2509(79)85087-3)
- [63] Donaldson TL, Nguyen YN. Carbon dioxide reaction kinetics and transport in aqueous amine membranes. *Ind Eng Chem Fundam* 1980;19:260-6. <https://doi.org/10.1021/i160075a005>
- [64] Perinu C, Arstad B, Jens KJ. NMR spectroscopy applied to amine–CO₂–H₂O systems relevant for post-combustion CO₂ capture: A review. *Int J Greenh Gas Control* 2014;20:230-43. <https://doi.org/10.1016/j.ijggc.2013.10.029>
- [65] Zhang R, Jiang W, Liang Z, Luo X, Yang Q. Study of equilibrium solubility, heat of absorption, and speciation of CO₂ absorption into aqueous 2-methylpiperazine (2MPZ) solution. *Ind Eng Chem Res* 2018;57:17496-503. <https://doi.org/10.1021/acs.iecr.8b03104>
- [66] Zheng W, Yan Z, Zhang R, Jiang W, Luo X, Liang Z, Yang Q, Yu H. A study of kinetics, equilibrium solubility, speciation and thermodynamics of CO₂ absorption into benzylamine (BZA) solution. *Chem Eng Sci*. 2022;251:117452. <https://doi.org/10.1016/j.ces.2022.117452>

List of Publications

Patent filed:

Monoj Kumar Mondal and **Ashish Gautam**. An aqueous amine blend for post-combustion capture (**Application number: 202411034454**).

Publications in Science Citation Indexed (SCI) Journals:

1. **Ashish Gautam** and Monoj Kumar Mondal. Post-combustion CO₂ absorption-desorption performance of novel aqueous binary amine blend of Hexamethylenediamine (HMDA) and 2-Dimethylaminoethanol (DMAE). *Energy (Elsevier Publication)* **2024**; 296: 130982. (**Impact factor = 9**)
2. **Ashish Gautam** and Monoj Kumar Mondal. Novel aqueous amine blend of 2-(Butylamino) ethanol and 2-Dimethylaminoethanol for CO₂ capture: Equilibrium CO₂ loading, RSM optimization, desorption study, characterization and toxicity assessment. *Separation and Purification Technology (Elsevier Publication)* **2023**; 322: 124279. (**Impact factor = 8.1**)
3. **Ashish Gautam** and Monoj Kumar Mondal. Post-combustion capture of CO₂ using novel aqueous Triethylenetetramine and 2-Dimethylaminoethanol amine blend: equilibrium CO₂ loading-empirical model and optimization, CO₂ desorption, absorption heat, and ¹³C NMR analysis. *Fuel (Elsevier Publication)* **2023**; 331: 125864. (**Impact factor = 6.7**)
4. **Ashish Gautam** and Monoj Kumar Mondal. Review of recent trends and various techniques for CO₂ capture: Special emphasis on biphasic amine solvents. *Fuel (Elsevier Publication)* **2023**; 334: 126616. (**Impact factor = 6.7**)

-
-
5. Akhil Kumar Gupta, **Ashish Gautam**, and Monoj Kumar Mondal. Experimental, modeling and RSM optimization of CO₂ loading for an aqueous blend of Diethylenetriamine and 3-Dimethyl amino-1-propanol. *Korean Journal of Chemical Engineering (Springer Publication)* **2023**; 40(5): 1151-67. (**Impact factor = 2.9**)

Book chapters:

1. **Ashish Gautam** and Monoj Kumar Mondal. Chapter: Recent advances in Microalgae process for post-combustion CO₂ capture (**Book**: Microbial Niche Nexus Sustaining Environmental Biological Wastewater and Water-Energy-Environment Nexus – By *Springer*) (**Accepted**).
2. **Ashish Gautam** and Monoj Kumar Mondal. Chapter 16: Algal bioplastics/ Polyhydroxybutyrate production: algal diversity, biosynthetic pathways, extraction methods and technical aspects, current prospects/applications, market trends (**Book**: Algae-derived Biochemicals of Industrial Importance – By *Springer Cham*) (**Accepted**).
3. **Ashish Gautam** and Monoj Kumar Mondal. Chapter 10: Scope of Food Waste to Resource Recovery: A Way of Green Advocacy (**Book**: Resource Recycling and Management of Food Waste – By *Springer Nature*) (**Communicated for publication**).

International conferences:

1. **Ashish Gautam** and Monoj Kumar Mondal. 2-(Methylamino)ethanol and 2-Dimethylaminoethanol aqueous amine blend for post-combustion CO₂ capture– Targeting net-zero emission. **IICHe-CHEMCON**, organized by **Heritage Institute of Technology, Kolkata** during 27-30 December, 2023.

2. **Ashish Gautam** and Monoj Kumar Mondal. Aqueous amine blend of 1,4-Butanediamine and 2-Dimethylaminoethanol for post-combustion CO₂ capture: An approach towards carbon neutrality. **ICPHD**, organized by **IIT Guwahati** during 03-05 November, 2023.
3. **Ashish Gautam** and Monoj Kumar Mondal. Experimental investigation of post-combustion CO₂ capture by aqueous amine blend of 2-(2-Aminoethylamino)ethanol and 2-Dimethylaminoethanol. **TISD**, organized by **MNNIT Allahabad, Prayagraj** during 27-29 October, 2023. (**Best oral presentation award**)
4. **Ashish Gautam** and Monoj Kumar Mondal. Aqueous Piperazine (PZ) and 2-Dimethylaminoethanol (DMAE) amine blend for post-combustion CO₂ capture: Absorption and desorption study. **CHEM-TECHNOVA**, organized by **HBTU, Kanpur** during 26 to 27 May, 2023.
5. **Ashish Gautam** and Monoj Kumar Mondal. Aqueous Triethylenetetramine and 3-Dimethylamino-1-propanol amine blend for post-combustion CO₂ capture. **CHEMCON**, organized by **HBTU, Kanpur** during 27-30 December, 2022. (**Best oral presentation award**)
6. **Ashish Gautam** and Monoj Kumar Mondal. Triethylenetetramine and 2-(Diethylamine)ethanol aqueous amine blend for Post-combustion CO₂ capture. **CHEMSMART**, organized by **NIT, Rourkela** during 16-18 December, 2022.
7. **Ashish Gautam** and Monoj Kumar Mondal. 1,5-Diamino-2-methylpentane and 2-Dimethylaminoethanol aqueous amine blend for post-combustion CO₂ capture. **ABC-HEE**, organized by **MNNIT Allahabad, Prayagraj** during 20-22 October, 2022.

8. **Ashish Gautam** and Monoj Kumar Mondal. Aqueous Triethylenetetramine and 2-Dimethylaminoethanol amine blend for post-combustion CO₂ capture: Equilibrium CO₂ loading, desorption study, empirical modeling and ¹³C NMR analysis. **CHEMTSF**, organized by **IIT, Roorkee** during 8-10 September, 2022.

Workshops and Training programs:

1. National Training Program on “Water Urbanism for Resilient Development: Case Studies from Varanasi, India” organized by **IIT (BHU)** and **National Institute of Disaster Management (NIDM)** New Delhi, (**Ministry of Home Affairs, Govt. of India**) on 04-06 September, 2023.
2. **LaTeX** - Introductory workshop organized by Teaching and Learning cell, **IIT (BHU)**, **Varanasi**, India on 15 April, 2023.